

Instructions for Parts Books

The heading at the top of the page will be the name for the picture of the parts, as it is for the page with the part numbers.

Definitions of various headings on part number pages:

Ref No. → The number referring to the part on the picture page of the parts book.

Part No. → The number assigned to this part at the time of printing the parts book.

Description → Description of part.

Qty. → Quantity of part used on that page.

S.N. → Serial number range of the equipment that this part will fit.
 → Your serial number should be greater than or equal to the number shown.
 → Your serial number should be lower than or equal to the serial number shown.

IC → Interchangeability of part shown.
 → The part shown will work for the older and later serial numbers.
 → The part shown is not interchangeable with the other serial numbers.
 → new for old
 → old for new

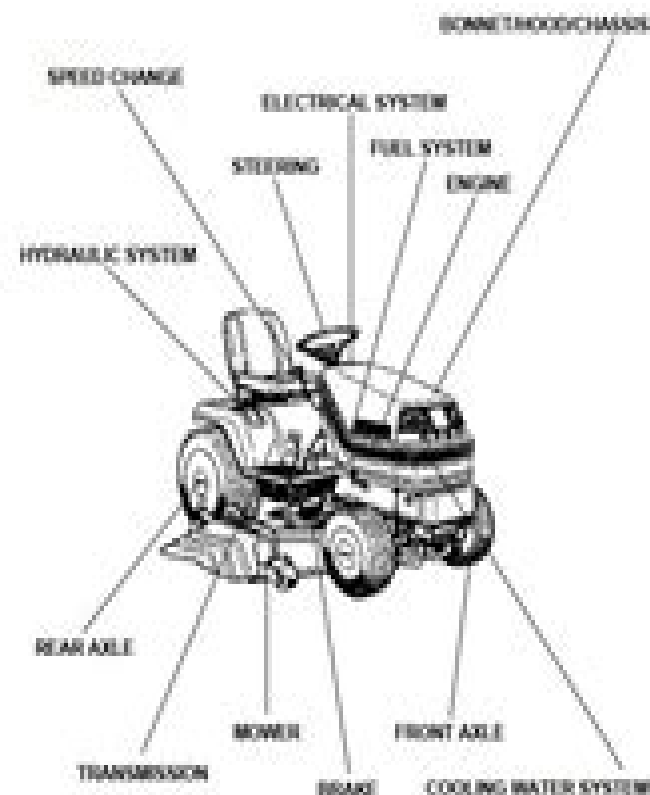
Remarks → Notes about this particular part. Sometimes this may say except (FR) or (CA).

Note about Country Codes:

The parts books we use in the United States of America may also contain information about a part that would be different in another country. It is important that you are aware of the country codes in some instances.

Symbol	Country
A	United States of America
CA	Canada
FR	France
GR	Greece
PO	Portugal
AO	Australia

If the remarks column says for example EXCEPT (FR) this part works in all countries except France. That means it will work here in the USA.



Kubota G2000 Owners Manual

AN Whitehead



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$$m = \frac{M}{1000d}$$

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web 1. molarity $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$ $M = \frac{0.4500 \text{ kg}}{0.0129 \text{ m}^3 \times 141.9579 \text{ g/mol}} = 0.825 \text{ mol/L}$
 0.00581158 mol / 0.00581158 mol = 1.0000
 0.4500 kg / 0.0129 m³ = 34.8837 mol/m³
 34.8837 mol/m³ / 141.9579 g/mol = 0.2458 mol/L

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web apr 7 2014 molarity moles of solute / liters of solution
 8.422 a first convert 250 ml to liters 250 / 1000 = 0.25 then
 calculate molarity 5 moles / 0.25 liters = 20 M
 3. c a solution with molarity 2 requires 2 moles of NaOH per liter so 4 x 2.8 M = 11.2 M
 a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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 $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$
 $M_{\text{mix}} = \frac{1 \times 480 + 2 \times 520 + 3 \times 720 + 4 \times 624 + 5 \times 1000}{1344} = 2.5 \text{ M}$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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 2. what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution
 3. 140 g of KCl is dissolved in 600 ml of water what is the molarity
 4. 724.4 g of ammonium phosphate in 4500 ml of alcohol

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 0.444 mol of COCl₂ in 0.654 L of solution
 98.0 g of phosphoric acid H₃PO₄ in 1.00 L of solution
 0.2074 g of calcium hydroxide Ca(OH)₂ in 40.00 ml of solution
 10.5 kg

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