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web equilibrium moles  $\text{hscn}$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{hscn}$  knowing the number of moles of  $\text{fe}^{3+}$  and  $\text{hscn}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two

species mol Fe 3 1 70 x 10<sup>5</sup> mol Fe 3 8 50 x 10<sup>4</sup> m

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