

# *Hart's Choice*

A Novella



CAROLYN ARNOLD

# Harts Choice A Novella

**J Spring**



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web equilibrium moles  $\text{HSCN}$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{HSCN}$  knowing the number of moles of  $\text{Fe}^{3+}$  and  $\text{HSCN}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $[\text{Fe}^{3+}] = 3.170 \times 10^{-5} \text{ mol/L}$   $[\text{Fe}^{3+}] = 3.850 \times 10^{-4} \text{ M}$

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