

## LAB MANUAL WORK INSTRUCTION

**Title of Experiment:** Reaction of Baking Soda and Vinegar

**Objective:** To observe the chemical reaction between baking soda (sodium bicarbonate) and vinegar (acetic acid) and to demonstrate the production of carbon dioxide gas.

### Pre-Lab Questions

- Write the chemical formula for sodium bicarbonate (baking soda) and acetic acid (vinegar).
- Predict what you think will happen when these two substances are mixed. Why?
- What gas do you expect to be produced from this reaction?

### Safety Precautions

- Wear safety goggles and gloves to protect against splashes.
- Perform the experiment in a well-ventilated area to avoid irritation of gases.
- Avoid direct skin contact with vinegar as it is an acid, though mild.

### Equipment and Materials

- |                                  |                   |
|----------------------------------|-------------------|
| 100-ml Beaker                    | Balloon           |
| 10-ml Graduated Cylinder         | Funnel            |
| Baking Soda (Sodium Bicarbonate) | Safety Goggles    |
| Vinegar (Acetic Acid)            | Laboratory Gloves |

### Procedure

- Place the beaker on a flat surface.
- Using the funnel, fill the balloon with two scoops of baking soda.
- Measure 50 ml of vinegar using the graduated cylinder and pour it into the beaker.
- Carefully fit the neck of the balloon over the mouth of the beaker without letting the baking soda fall into the beaker.
- Once the balloon is securely in place, lift the balloon allowing the baking soda to fall into the vinegar.
- Observe the reaction between the baking soda and vinegar.
- Note the speed and amount of gas produced as indicated by the inflation of the balloon.
- Record any changes in temperature of the beaker.
- Carefully remove the balloon from the beaker.
- Dispose of the vinegar and baking soda solution in the sink and reuse the beaker.

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# Lab Manual Format

**James G. Cappuccino, Natalie Sherman**



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