

# Acids and bases

## Chapter outline

- Recall the definitions of Brønsted–Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid–base pairs.
- Recall the terms ‘amphoteric’ and ‘amphiprotic’.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction. **HL**
- Deduce the dissociation constant for weak acids and bases. **HL**
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the  $pK_a$  or  $pK_b$  value of a weak acid or base. **HL**
- Explain how an acid–base indicator works and select a suitable indicator for a titration. **HL**

## KEY TERMS

**Acid deposition:** A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth’s surface.

**Brønsted–Lowry acid:** A proton ( $H^+$ ) donor.

**Brønsted–Lowry base:** A proton ( $H^+$ ) acceptor.

**pH:** A measure of the concentration of hydrogen ions in an aqueous solution.  
 $pH = -\log_{10}[H^+]$

**Strong acid/base:** An acid/base that is completely dissociated into its ions in aqueous solution.

**Weak acid/base:** An acid/base that is only partially dissociated into its ions in aqueous solution.

# Guided Reading And Study Workbook Chapter 20 Acids And Bases

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## **Table of Contents Guided Reading And Study Workbook Chapter 20 Acids And Bases**

1. Understanding the eBook Guided Reading And Study Workbook Chapter 20 Acids And Bases
  - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 20 Acids And Bases
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 20 Acids And Bases
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Reading And Study Workbook Chapter 20 Acids And Bases
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 20 Acids And Bases
  - Personalized Recommendations
  - Guided Reading And Study Workbook Chapter 20 Acids And Bases User Reviews and Ratings
  - Guided Reading And Study Workbook Chapter 20 Acids And Bases and Bestseller Lists
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  - Guided Reading And Study Workbook Chapter 20 Acids And Bases eBook Subscription Services
  - Guided Reading And Study Workbook Chapter 20 Acids And Bases Budget-Friendly Options
6. Navigating Guided Reading And Study Workbook Chapter 20 Acids And Bases eBook Formats
    - ePub, PDF, MOBI, and More
    - Guided Reading And Study Workbook Chapter 20 Acids And Bases Compatibility with Devices
    - Guided Reading And Study Workbook Chapter 20 Acids And Bases Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Interactive Elements Guided Reading And Study Workbook Chapter 20 Acids And Bases
  8. Staying Engaged with Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Guided Reading And Study Workbook Chapter 20 Acids And Bases
  9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 20 Acids And Bases
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Setting Reading Goals Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 20 Acids And Bases
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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