

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Lab Section: \_\_\_\_\_

# Experiment 2: Molecular Models

## REPORT SHEET

## Pre-lab Study Questions:

- Objectives of the experiment (in your own words):
- What kinds of electrons form the covalent bonds in molecules?
- Is a polar molecule symmetrical or asymmetrical?

FORMULA OF MOLECULE	LEWIS STRUCTURE	3-D SHAPE w/ lone pairs	MOLECULAR GEOMETRY	MOLECULAR POLARITY
example: CH <sub>4</sub>	$  \begin{array}{c}  \text{H} \\    \\  \text{H}-\text{C}-\text{H} \\    \\  \text{H}  \end{array}  $		tetrahedral	nonpolar
H <sub>2</sub> O	$  \begin{array}{c}  \text{H} \quad \text{H} \\  \diagdown \quad \diagup \\  \text{O} \\  \diagup \quad \diagdown \\  \text{H} \quad \text{H}  \end{array}  $	$  \begin{array}{c}  \text{H} \quad \text{H} \\  \diagdown \quad \diagup \\  \text{O} \\  \diagup \quad \diagdown \\  \text{H} \quad \text{H}  \end{array}  $	tetrahedral	polar
H <sub>2</sub> O <sup>+</sup>	$  \begin{array}{c}  \text{H} \quad \text{H} \\  \diagdown \quad \diagup \\  \text{O} \\  \diagup \quad \diagdown \\  \text{H} \quad \text{H}  \end{array}  $	$  \begin{array}{c}  \text{H} \quad \text{H} \\  \diagdown \quad \diagup \\  \text{O} \\  \diagup \quad \diagdown \\  \text{H} \quad \text{H}  \end{array}  $	tetrahedral	nonpolar
C <sub>2</sub> H <sub>6</sub>				
C <sub>2</sub> H <sub>4</sub>	$  \begin{array}{c}  \text{H} \quad \text{H} \\  \diagdown \quad \diagup \\  \text{C}=\text{C} \\  \diagup \quad \diagdown \\  \text{H} \quad \text{H}  \end{array}  $		trigonal planar	non polar
C <sub>2</sub> H <sub>2</sub>				
CO <sub>2</sub>	$  \begin{array}{c}  \text{O}=\text{C}=\text{O}  \end{array}  $		Linear	Non polar
H <sub>2</sub> O <sub>2</sub>				
CH <sub>3</sub> Cl				

## Lab 22 Models Molecular Compounds Answer

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Development Administration.  
Technical Information Center**



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**Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

Polarity, Solutions, and Separation Science Kenda Jo Lemont, 1998      Selected Water Resources Abstracts , 1988-12

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