

Experiment 15: Specific Heat of a Metal

Purpose: To determine the specific heat of a substance.

Procedure:

Record all data in Data Table 1.

1. Heat 250 mL of water in a 400-mL beaker until it is boiling gently.
2. While the water is heating, determine and record the mass of a clean, dry 50-mL beaker to the nearest 0.01 g. Add between 80 g and 120 g of lead shot to the beaker. Measure the combined mass of the beaker and lead to the nearest 0.01 g and record the measurement.
3. Transfer the lead shot to a large, dry test tube. Use the utility clamp to suspend the test tube in the boiling water as shown in Figure 15.1. The lead shot should be below the level of the water in the beaker. Leave the test tube in the boiling water bath for at least 10 minutes.
4. While the lead shot is heating, measure 100 mL of distilled water in a graduated cylinder. Pour the water into a Styrofoam cup and place the cup in a 250-mL beaker for support.
5. Measure and record the temperature of the water in the Styrofoam cup and of the water in the boiling bath.
6. Remove the test tube from the boiling water and quickly pour the lead shot into the water-filled, plastic-foam cup. Place a thermometer and a glass stirring rod into the cup. Use the stirring rod to gently stir the lead shot. *Do not stir the shot with the thermometer.* Note the temperature frequently and record the maximum temperature reached.
7. Pour the water off and return the lead shot to your teacher.
8. Follow the same procedure to determine the specific heat of a second metal.

Data Table 1:

Measurements of Mass and Temperature		
	Trial 1 (Pb)	Trial 2 (Al)
Mass of metal shot (g)	55.7	68.1
Mass of water in cup (g)	100.0	100.0
Initial temperature of water in cup ($^{\circ}\text{C}$)	22.0	23.0
Initial temperature of metal shot ($^{\circ}\text{C}$) (same as temperature of boiling water)	101.1	99.8
Maximum temperature of metal + water	27.1	32.0

Analysis Questions:

1. Determine the changes in temperature of the water (ΔT_{water}) and of the lead shot (ΔT_{metal}) for each trial.
2. Calculate the heat gained by the water in each trial.

Lab 22 Specific Heat Of Metals Answers

**University of the State of New York.
High School Department**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the text box.

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