

electrochemistry - the study of electricity - related applications of oxidation - reduction reactions

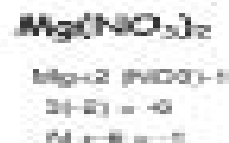
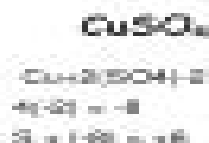
oxidation numbers -

- numbers assigned to the elements in a chemical reaction that give the apparent charge of the elements
- In ionic compounds - oxidation # = true charge
- In molecular compounds - oxidation # is based on electron distribution
- Oxidation #s are written with the sign first at and then the number

Rules For Assigning Oxidation Numbers:  
situation

|                                |  | oxidation #   |  |
|--------------------------------|--|---------------|--|
| <u>uncombined element</u>      |  | 0             |  |
| <u>sum in a compound</u>       |  | 0             |  |
| <u>sum in a polyatomic ion</u> |  | charge on ion |  |
| <u>elements in a compound:</u> |  |               |  |
| Fluorine                       |  | -1            |  |
| Oxygen                         |  | -2            |  |
| Hydrogen                       |  | +1            |  |
| group 1                        |  | +1            |  |
| group 2                        |  | +2            |  |
| Aluminum                       |  | +3            |  |

Example:



# Gpb Note Taking Guide 1401 Answer Key

**Michael Hoelscher**



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