

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$$Rhc = A$$

$$A = 2.18 \times 10^{-18} \text{ J}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?

$$\lambda = \frac{c}{\nu}$$

- a) 0.307 m d) 0.110 m

- b) 0.0365 m c) 27.4 m

- c) 0.122 m $\frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?

- a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$

- b) 3.30 m d) 2.73 m

$$\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$$

3. Which one of the following is NOT a proper unit for frequency?

- a) Hz

- b) s^{-1}

- c) $\text{m}\cdot\text{s}^{-1}$

- d) $\frac{1}{\text{sec}}$

velocity

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.

$$n=6 \rightarrow n=2$$

$$E_6 = -2.18 \times 10^{-18} \text{ J} - E_2 = -2.18 \times 10^{-18} \text{ J}$$

- a) 0.12334 m

- d) $4.1029 \times 10^{-7} \text{ m}$

- b) 24.373 m

- e) 36.559 m

- c) $2.7353 \times 10^{-7} \text{ m}$

$$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} = 4.1 \times 10^{-7} \text{ m}$$

$$\Delta E = 4.84 \times 10^{-19} \text{ J}$$

5. What is the relationship between the energy of a photon of light and its frequency?

- a) $E = \nu$

- d) $E = \frac{1}{h\nu}$

- b) $E = \frac{h}{\nu}$

- e) $E = \frac{\nu}{h}$

- c) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?

$$\Delta E = E_3 - E_2$$

- a) $1.52 \times 10^4 \text{ J}$

- d) $4.48 \times 10^{-19} \text{ J}$

- b) $3.63 \times 10^{-19} \text{ J}$

- e) $3.03 \times 10^{-19} \text{ J}$

- c) $2.18 \times 10^{-19} \text{ J}$

$$-2.18 \times 10^{-18} \text{ J} - (-2.18 \times 10^{-18} \text{ J}) = 3.03 \times 10^{-19} \text{ J}$$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

- a) $3.03 \times 10^{-12} \text{ m}$

- c) $3.30 \times 10^{11} \text{ m}$

- b) $2.42 \times 10^{-12} \text{ m}$

- d) $1.59 \times 10^{-25} \text{ m}$

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{9.11 \times 10^{-31} \text{ kg} \times 0.8 \times 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}} = 3.03 \times 10^{-12} \text{ m}$$

8. What resultant is expected from the interference of the two waves shown below?



- a)

- c)

- b)

- d)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

- a) n

- c) m_l

- b) ℓ

- d) m_s

Gpb Chemistry Answers Atomic Structure

Michal Rosen-Zvi



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Mechanisms of Ion Formation and Fragmentation by Fast-atom Bombardment and Matrix-assisted Laser Desorption Ionization Mass Spectrometry Gabriella Székely,1997 **U.S. Government Research Reports** ,1954
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periodic table and how it organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element's position on the table In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry So let's dive in and unlock the mysteries of the atom together

Chemistry and Atomic Structure Jd Main Smith,2023-07-18 This book is an introduction to the chemistry of atomic structure It covers the basic principles of atomic theory including the nature of atoms chemical bonds and chemical reactions The author provides clear explanations of complex concepts making this book accessible to students and non experts alike Anyone interested in chemistry or atomic theory will find this book a useful resource This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant

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Table of Contents **Gpb Chemistry Answers Atomic Structure**

1. Understanding the eBook **Gpb Chemistry Answers Atomic Structure**
 - The Rise of Digital Reading **Gpb Chemistry Answers Atomic Structure**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Gpb Chemistry Answers Atomic Structure**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Gpb Chemistry Answers Atomic Structure**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Gpb Chemistry Answers Atomic Structure**
 - Personalized Recommendations
 - **Gpb Chemistry Answers Atomic Structure** User Reviews and Ratings
 - **Gpb Chemistry Answers Atomic Structure** and Bestseller Lists

5. Accessing Gpb Chemistry Answers Atomic Structure Free and Paid eBooks
 - Gpb Chemistry Answers Atomic Structure Public Domain eBooks
 - Gpb Chemistry Answers Atomic Structure eBook Subscription Services
 - Gpb Chemistry Answers Atomic Structure Budget-Friendly Options
6. Navigating Gpb Chemistry Answers Atomic Structure eBook Formats
 - ePub, PDF, MOBI, and More
 - Gpb Chemistry Answers Atomic Structure Compatibility with Devices
 - Gpb Chemistry Answers Atomic Structure Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Chemistry Answers Atomic Structure
 - Highlighting and Note-Taking Gpb Chemistry Answers Atomic Structure
 - Interactive Elements Gpb Chemistry Answers Atomic Structure
8. Staying Engaged with Gpb Chemistry Answers Atomic Structure
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Chemistry Answers Atomic Structure
9. Balancing eBooks and Physical Books Gpb Chemistry Answers Atomic Structure
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Chemistry Answers Atomic Structure
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gpb Chemistry Answers Atomic Structure
 - Setting Reading Goals Gpb Chemistry Answers Atomic Structure
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gpb Chemistry Answers Atomic Structure
 - Fact-Checking eBook Content of Gpb Chemistry Answers Atomic Structure
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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