

**THE UNIVERSITY OF CHICAGO**  
**DEPARTMENT OF CHEMISTRY**  
**RESEARCH REPORT**

**RESEARCH REPORT**

1. **Introduction**  
The purpose of this report is to describe the results of the experiments conducted during the summer of 1964. The experiments were designed to study the effect of temperature on the rate of reaction between hydrogen peroxide and potassium permanganate in the presence of ceric sulfate as a catalyst. The reaction was studied at various temperatures ranging from 10°C to 40°C. The rate of reaction was determined by measuring the volume of oxygen gas evolved over a period of time. The results show that the rate of reaction increases with increasing temperature. The activation energy of the reaction was calculated to be 45 kJ/mol.

2. **Experimental Procedure**  
The reaction was carried out in a 250 ml Erlenmeyer flask. A solution of potassium permanganate (0.01 M) was added to a solution of hydrogen peroxide (0.1 M) in the presence of ceric sulfate (0.001 M) as a catalyst. The reaction was initiated by the addition of a small amount of sulfuric acid. The volume of oxygen gas evolved was measured by the displacement of water in a graduated cylinder. The temperature of the reaction mixture was maintained at a constant value by using a water bath. The rate of reaction was determined by measuring the volume of oxygen gas evolved over a period of time. The activation energy of the reaction was calculated from the Arrhenius plot.

# Mcculloch Service Manual Si 136

**Michael Hoelscher**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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