Grade 10 Chemistry Review

- Use the periodic table to answer the following questions concerning calcium.
 - a) In what period is it located?
 - b) In what group is it located?
 - c) What is its atomic number?
 - d) What is its atomic mass?
 - e) In what state would it be found at room temperature?
- Explain why O forms a -2 charged ion. Provide a written explanation and Bohr-Rutherford diagrams for both the atom and the ion.
- Write formulas for each of the following compounds.

sulfur pentafluoride calcium oxide nitrogen gas iron (III) sulfide sodium sulfate tin(II) bromide calcium phosphate ammonia

Name the type of reaction and balance each equation.

 Al_2O_3 a) FeO Type: b) Ca Ca₃P₂ Type: c) KBr + La -0e KI Brz. Type: d) CaHa O_2 CO_2 H_2O type: e) Ba(OH)₂ HCI -3-BaCl₂ H₂O type: f) SnO Sn. Type: -30 O_{2}

- Write word equations for each of the reactions in #4. C₃H₈ is propane.
- 6. Describe a standard test for each of the following substances:
 - a) oxygen gas
 - b) hydrogen gas
 - c) carbon dioxide gas
 - d) water vapour

Grade 10 Science Chemistry Unit Test Cscope

Ken Dobson

Grade 10 Science Chemistry Unit Test Cscope:

Co-ordinated Science Ken Dobson,1993

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