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Reihe: Lehrerfortbildung

Kinder begreifen Mathematik

Frühe mathematische Bildung
und Förderung



1. Auflage

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 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{CO}_2(\text{g}) + 3 \text{H}_2(\text{g}) + \text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of $NaOH$ combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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