

## Post-Lab Questions

1. From the experiment, which reactant is the limiting reactant? Which one is the reactant in excess? How do you determine which one is the limiting reactant?
2. If you weigh the watch glass then write your name on it with the wax pencil, how will that affect your percent yield?
3. What factors could cause the percent yield of a product to be less than 100%?
4. What factors could cause the percent yield of a product to be more than 100%?
5. Balance the following reaction:  $\text{___ C}_3\text{H}_8 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$ 
  - a. 4.30 g of  $\text{C}_3\text{H}_8$  reacts with 4.30 g of  $\text{O}_2$ . Which one is the limiting reagent, and which one is in excess? Explain your reasoning.
  - b. Based on the limiting reagent, what are the theoretical yields of  $\text{CO}_2$  and  $\text{H}_2\text{O}$ ?
  - c. You perform the experiment and produce 2.67 g of  $\text{CO}_2$  and 0.79 g of  $\text{H}_2\text{O}$ . What are your percent yields for each product?

# Gpc Chem 1211 Lab Manual

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