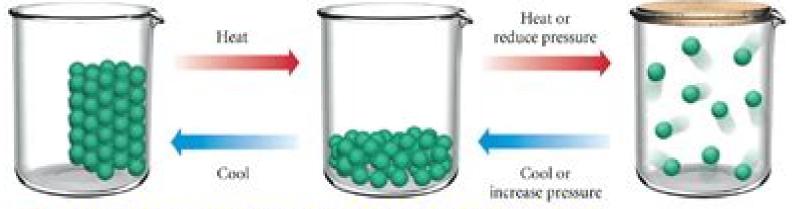
# Chapter 11: Liquids and Solids

### A. Intermolecular Forces

- The addition of heat changes the state of a compound from solid to liquid to gas as kinetic and potential energy are added to the molecules
  - a. Kinetic Energy involves the motion of the molecules
  - Potential Energy involves how far apart the molecules are



- Intermolecular Forces hold the molecules together
  - a. Ion-Dipole Interactions: Na\*Cl interactions with water
  - b. Dipole-Dipole Attractions: interaction of polar molecules
  - e. Hydrogen Bonding: special dipole-dipole attraction in molecules where H is bonded to an electronegative atom (O, N, F, etc...)
    - Small size of H lets the dipoles get very close
    - ii. Large electronegativity difference makes attraction great

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