

Heat of Neutralization

Introduction:

The experiment was used to determine the heat evolved in an aqueous acid-base reaction. For strong acids and bases reacting in sufficiently dilute solutions, the heat evolved is approximately 57 kJ mol^{-1} at 298 K . Increasing the concentration should increase the heat of the reaction, while using a weaker acid or base will reduce the heat of the reaction. A series of acid-base neutralizations were performed on aqueous sodium hydroxide and the heat of the reaction was determined calorimetrically.

Experimental:

To measure the heat capacity of the calorimeter system, 225 mL of distilled water was added to a clean, dry Dewar. The water was stirred and the temperature read every 20 seconds until it became constant. Then, after recording the temperature of it, 25 mL of ice water was added to the flask. The mixture was stirred and the temperature was read until constant readings were obtained. The Dewar flask was then dried and 200 mL of the 0.2 N NaOH solution was added to the calorimeter. The temperature was taken before 50 mL of the 0.8 N HCl solution was added. The mixture was stirred and the temperature was read until constant readings were obtained. This process was then repeated twice for both the sulfuric and phosphoric acid solutions.

Results:

Calorimeter				Heat Capacity of Calorimeter	
Q	T _i	T _f	ΔT	$q = m \cdot c \cdot \Delta T$	$C = \frac{q}{\Delta T}$
$+10.25$	20.0°C	18.0°C	-2.0	$100 \cdot 4.18 \cdot (-2.0)$	$C = \frac{-836 \text{ J}}{-2.0^\circ\text{C}} = 418 \text{ J/}^\circ\text{C}$
$+20.50$	20.0°C	18.0°C	-2.0	$200 \cdot 4.18 \cdot (-2.0)$	$C = \frac{-1672 \text{ J}}{-2.0^\circ\text{C}} = 836 \text{ J/}^\circ\text{C}$
$+20.50$	20.0°C	18.0°C	-2.0	$200 \cdot 4.18 \cdot (-2.0)$	$C = \frac{-1672 \text{ J}}{-2.0^\circ\text{C}} = 836 \text{ J/}^\circ\text{C}$
HCl (0.8 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$q = m \cdot c \cdot \Delta T$	$\Delta H = \frac{q}{n}$
$+10.25$	20.0°C	21.0°C	$+1.0$	$100 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{418 \text{ J}}{0.1 \text{ mol}} = 4180 \text{ J/mol}$
$+20.50$	20.0°C	21.0°C	$+1.0$	$200 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{836 \text{ J}}{0.1 \text{ mol}} = 8360 \text{ J/mol}$
$+20.50$	20.0°C	21.0°C	$+1.0$	$200 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{836 \text{ J}}{0.1 \text{ mol}} = 8360 \text{ J/mol}$
H ₂ SO ₄ (0.4 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$q = m \cdot c \cdot \Delta T$	$\Delta H = \frac{q}{n}$
$+10.25$	20.0°C	21.0°C	$+1.0$	$100 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{418 \text{ J}}{0.05 \text{ mol}} = 8360 \text{ J/mol}$
$+20.50$	20.0°C	21.0°C	$+1.0$	$200 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{836 \text{ J}}{0.05 \text{ mol}} = 16720 \text{ J/mol}$
$+20.50$	20.0°C	21.0°C	$+1.0$	$200 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{836 \text{ J}}{0.05 \text{ mol}} = 16720 \text{ J/mol}$
H ₃ PO ₄ (0.3 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$q = m \cdot c \cdot \Delta T$	$\Delta H = \frac{q}{n}$
$+10.25$	20.0°C	21.0°C	$+1.0$	$100 \cdot 4.18 \cdot (+1.0)$	$\Delta H = \frac{418 \text{ J}}{0.1 \text{ mol}} = 4180 \text{ J/mol}$
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Picture 1: Data tables and calculations for change in temperature and heat of neutralization

Heat Of Neutralization Lab Answer Key

WJ Hussar



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A Study of Methods for Determining Heats of Neutralization at Different Temperatures John Bernard Dickson,1917

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Table of Contents Heat Of Neutralization Lab Answer Key

1. Understanding the eBook Heat Of Neutralization Lab Answer Key
 - The Rise of Digital Reading Heat Of Neutralization Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Heat Of Neutralization Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heat Of Neutralization Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heat Of Neutralization Lab Answer Key
 - Personalized Recommendations
 - Heat Of Neutralization Lab Answer Key User Reviews and Ratings
 - Heat Of Neutralization Lab Answer Key and Bestseller Lists
5. Accessing Heat Of Neutralization Lab Answer Key Free and Paid eBooks
 - Heat Of Neutralization Lab Answer Key Public Domain eBooks
 - Heat Of Neutralization Lab Answer Key eBook Subscription Services
 - Heat Of Neutralization Lab Answer Key Budget-Friendly Options
6. Navigating Heat Of Neutralization Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Heat Of Neutralization Lab Answer Key Compatibility with Devices
 - Heat Of Neutralization Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Heat Of Neutralization Lab Answer Key
 - Highlighting and Note-Taking Heat Of Neutralization Lab Answer Key
 - Interactive Elements Heat Of Neutralization Lab Answer Key

8. Staying Engaged with Heat Of Neutralization Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heat Of Neutralization Lab Answer Key
9. Balancing eBooks and Physical Books Heat Of Neutralization Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heat Of Neutralization Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Heat Of Neutralization Lab Answer Key
 - Setting Reading Goals Heat Of Neutralization Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Heat Of Neutralization Lab Answer Key
 - Fact-Checking eBook Content of Heat Of Neutralization Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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