

Calorimetry: Heat of Neutralization

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Purpose: Determine the calorimeter constant for a calorimeter. Determine the enthalpy of neutralization for the reaction of a strong acid and a strong base.

Background: When a system absorbs heat (q) its temperature changes by ΔT . Experimental measurements demonstrate that the heat absorbed by a system and its corresponding temperature change are directly proportional. $q = C \times \Delta T$. C is called the **heat capacity**, which is often defined as the quantity of heat required to change its temperature by 1°C .

$$C = q / \Delta T = \text{J} / ^\circ\text{C}$$

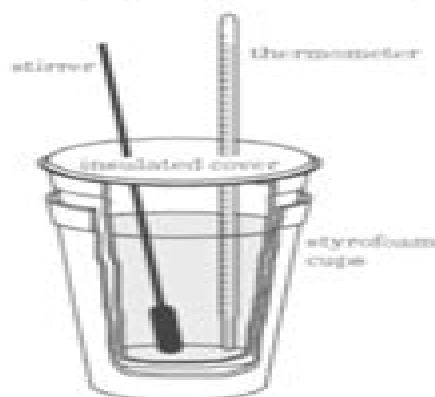
Heat capacity is an extensive property; it depends on the amount of matter being heated. The measure of intrinsic capacity of a substance to absorb heat is called **specific heat capacity** (C_s), the amount of heat required to raise the temperature of 1 gram of the substance by 1°C .

$$C_s = q / (\Delta T \times m)$$

The C_s of H_2O is $4.184 \text{ J/g } ^\circ\text{C}$. For the case where all the heat is used to heat water, the amount of heat is calculated by the relation

$$q = C_s \times m_{\text{water}} \times \Delta T$$

The enthalpy change for a chemical reaction, ΔH_{rxn} is also called the **enthalpy of reaction** or **heat of reaction**, which is also an extensive property. For many aqueous reactions, ΔH_{rxn} can be measured simply using a coffee-cup calorimeter.



Heat Of Neutralization Lab Manual

Sebastian Brünink



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Chemistry Lab Manual Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar, Lab Manual **Hard Bound Lab Manual Chemistry** Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar, Lab Manuals

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Analysis section **Laboratory Manual to Accompany Introductory Chemistry** Uno Kask, Otto W. Nitz, Martin G. Ondrus, Tammy J. Melton, 1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rational for each experiment **Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi** Arthur D. Baker, 1990 **Laboratory Manual in Quantitative Chemical Analysis , Laboratory Manual for Fundamentals of Chemistry 3/E** Jo A. Beran, 1988 **Analytical and Inorganic Chemistry Laboratory Manual** Mr. Rohit Manglik, 2024-03-24 Provides practical experiments and procedures in analytical and inorganic chemistry to reinforce theoretical concepts *Manual of Molecular and Clinical Lab Immunology* Barbara Detrick, Robert G. Hamilton, James D. Folds, 2006-03-10 Introduces new material that reflects the significant advances and developments in the field of clinical laboratory immunology Provides a comprehensive and practical approach to the procedures underlying clinical immunology testing Emphasizes molecular techniques used in the field of laboratory immunology Updates existing chapters and adds significant new material detailing molecular techniques used in the field Presents guidelines for selecting the best procedures for specific situations and discusses alternative procedures Covers aspects of immunology related disciplines such as allergy autoimmune diseases cancers and transplantation immunology

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