

Name: _____

Data

1. mass of clean empty crucible after heating 15.4709g
2. mass of crucible with magnesium strip 15.5854g
3. mass of magnesium strip 0.1145g
4. mass of crucible and product (magnesium oxide) 15.6589g
5. mass of product 0.0735g

Calculations

Label the type of chemical reaction, balance the chemical equation and calculate the Theoretical Yield of Magnesium Oxide based upon your amount of starting material.

Balance the Equation: $2\text{Mg (s)} + \text{O}_2 \text{ (g)} \rightarrow 2\text{MgO (s)}$

Reaction Type: Redox

Show all calculations for the theoretical yield of magnesium oxide

$$\begin{aligned} \text{molar mass} &= \frac{80\text{g}}{2 \times 24 + 48} = \frac{25}{60} = .3125 \\ \frac{0.0735}{.3125} &= .2352 \end{aligned}$$

Theoretical Yield of Magnesium Oxide = _____

Actual Yield of Magnesium Oxide = _____

Calculation of the % Yield:

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