

Grade 9 Natural Science Exam: Term 4

Name _____

Date _____

Class _____

Time: 2 hours



Total: / 100

Instructions for learners

- The paper consists of **100** questions and **100** questions.

SECTION A (2 marks)

SECTION B (2 marks)

- Read the full question carefully, according to the numbering system used in the question paper.

- In the case of calculations, show ALL steps.

- Round answers to 100 decimal places, unless specified.

- Graph paper is provided for SECTION B (a).

- Write neatly and legibly.



Section A: Multiple choice questions (100 marks)

For the following questions, write the letter of the correct answer in the box.



What is the correct answer?



Section B: Short answer questions (100 marks)

1. (10 marks)

2. (10 marks)

3. (10 marks)



Section C: Long answer questions (100 marks)

1. (10 marks)

2. (10 marks)

3. (10 marks)

4. (10 marks)

Section D: Table completion (100 marks)

Section D
1. Compression
2. Tension
3. Transverse force
4. Magnetic force
5. Tension



Section E: Graphing and calculation questions (100 marks)

1. (10 marks)

2. (10 marks)

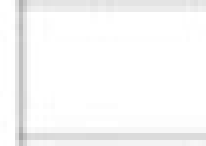
3. (10 marks)

4. (10 marks)

5. (10 marks)

6. (10 marks)

Total Section E: 60



7. (10 marks)

8. (10 marks)



Section F: Graphing and calculation questions (100 marks)

1. (10 marks)

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100. (10 marks)

Grade 9 Natural Science Midyear Exam

Kathleen Armour



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web an ion that is found in both salts in a solution common ion effect the lowering of the solubility of an ionic compound as a result of the addition of a common ion is called this lower the smaller the numerical value of the solubility product constant the the solubility of the compound precipitate

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web hexane is a nonpolar liquid with a dipole moment of zero and therefore does not significantly interact with the ions of the nacl crystals 11 a fe no 3 3 is a strong electrolyte thus it should completely dissociate into fe 3 and no 3 ions therefore z best represents the solution b fe no 3 3 s fe 3 a q 3 no 3 a q

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the ion product and the K_{sp} for that solution listed on the left the ion product exceeds

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web here the value of concentration of water is taken as 0 36 instead of 0 16 divide the product of the concentration of water carbon monoxide and the equilibrium constant by the concentration of hydrogen instead of multiplying it correct d the value of K_{eq} is taken as 4 4 instead of 0 44