ACID-BASE STUDY GUIDE

- Normal Plasma pH 7.35 to 7.45 (hydrogen (H+) concentration
 - o Higher this goes, higher the mortality rate
 - Measure the acidity or alkalinity of blood
 - o Homeostatic mechanisms keep pH normal
 - Buffer systems, kidneys & lungs
 - Greater H+ concentration = lower pH & more acidic solute
 - Lower H+ concentration = higher pH & more alkaline solute
- Major ECF buffer system
 - o Bicarbonate-carbonic acid system
 - CO2 +H2O ←→ H2CO3←→H+ and HCO3-
 - Kidneys regulate bicarbonate in the ECF
 - Lungs (controlled by medulla) regulate CO2, thus the carbonic acid in the ECF
 - Lungs are first responders
- Other buffer systems
 - ECF: inorganic phosphates and plasma proteins
 - ICF: proteins, organic, inorganic phosphates
 - Hemoglobin

METABOLIC ACIDOSIS

pH lower than 7.35

- Can occur by a gain of H+ or loss of bicarbonate ions (<22 mm Hg)
- Causes
 - Direct loss of bicarbonate (diarrhea, administration of IV nutrition w/o bicarb or bicarb-producing solutes, diuretic use, early renal insufficiency, excess chloride, lower intestinal fistulas, ureterostomies)
 - o Most common due to kidney injury
 - Lactic acidosis (higher it is, the worse sepsis is), salicylate poisoning, renal failure, methanol, ethylene, or propylene toxicity, DKA, ketoacidosis due to starvation
 - o Systemic issue: diabetes
- Clinical Manifestations
 - Headache, confusion, drowsiness, increased RR/depth, low BP, nausea, vomiting, cold/clammy skin, arrythmias, shock, peripheral vasodilation, decreased cardiac output
- Medical Management
 - First have to correct underlying issue

Key Acid Base Study Guide

Marcia L. Gillette,Ira Blei,Wendy Gloffke,George Odian



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