

Chemical Kinetics Lab Report

Title

Using Chemical Kinetics to Analyze Concentration and Reaction Rates

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Introduction

Chemical kinetics introduces the concept of understanding the rates of chemical reactions and, for the sake of this experiment, how the speed of molecular reconstruction is affected by an increase or decrease of a reactant's concentration. The rate of reaction quantitatively measures the change in the amount of reactant or product per unit of time (Flowers et al., 2015). To test this concept, this experiment analyzed the reaction between acetone (CH_3COCH_3) and iodine (I_2), represented by the following equation:

$CH_3COCH_3(aq) + I_2(aq) \rightarrow CH_3COCH_2I(aq) + H^+(aq) + I^-(aq)$. To enhance the significance of each chemical, the concentrations between acetone and iodine were altered amongst three trials and timed accordingly. The reaction was determined by the time the chemicals turned from yellow to colorless, which occurs due to the consumption of iodine as the reaction proceeds. By increasing an individual reactant and keeping the remaining chemicals constant during respected trials, it could be determined which concentration of a specific reactant held more of an effect on the reaction rate to find the relationship between reactants and reaction rates. Because acetone was known to have a greater surface area and evaporation rate compared to that of iodine, an initial claim was made that the fastest reaction rate would occur in the trial with an increase in the concentration of acetone.

Experimental Methods

Deionized water was introduced into the experiment with the knowledge that the liquid itself holds no effect on any chemical reaction. The total volume of each solution was to be measured out to 50mL; the addition of deionized water allowed for this experiment to avoid any inaccuracies. An additional liquid, hydrochloric acid, remained at a constant 10mL throughout the entirety of the experiment as it served as

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Valentine Korah



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