



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant (K_p):</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
<p>Equilibrium Constant, K_c</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
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<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>		
<p>Concentration</p>	<ul style="list-style-type: none">Increasing the concentration of a reactant will shift the equilibrium to the right.Increasing the concentration of a product will shift the equilibrium to the left.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 will shift the equilibrium to the right, producing more NH_3.</p>
<p>Pressure</p>	<ul style="list-style-type: none">Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas.Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure will shift the equilibrium to the right, producing more NH_3.</p>
<p>Temperature</p>	<ul style="list-style-type: none">Increasing the temperature will shift the equilibrium to the endothermic side.Decreasing the temperature will shift the equilibrium to the exothermic side.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature will shift the equilibrium to the left, producing less NH_3.</p>

Guide Chemical Equilibrium

Clemens Wendtner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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