

CH 4. RADICALS

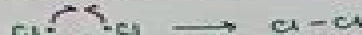
1. INITIATION - FORM RADICALS



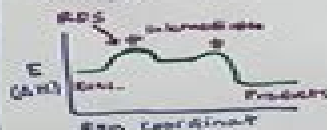
2. PROPAGATION - RADICAL REACTANT → RADICAL PRODUCT



3. TERMINATION - 2 RADICALS → NO RADICALS



$$\Delta H_{\text{rxn}} = \sum (\text{bonds broken}) - \sum (\text{bonds formed}) - \text{form bond}$$



F_2 - most reactive
 Cl_2 - reactive, not selective
 Br_2 - selective, picky
 I_2 - not reactive, endo.



RESONANCE + SUBST. ↑ RAD. STABILITY
 $\text{Allyl} \cdot > \text{R}-\text{C}(\text{H})-\text{R} > \text{R}-\text{C}(\text{H})_2-\text{R} > \text{R}-\text{C}(\text{H})_3$
 $3^\circ > 2^\circ > 1^\circ$

CH 5. STEREOCHEMISTRY

CHIRAL - Not symmetrical



ACHIRAL - SYMMETRY



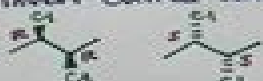
MESO - ACHIRAL W/ CHIRAL CENTERS



CONST. ISOMER - connect. diff.



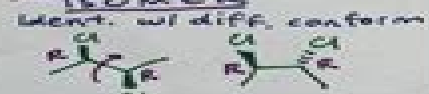
ENANTIOMER - Invert chiral centers



DIASTEROMER - Invert some, not all, chiral centers



CONFORMATIONAL ISOMER



CH 6. NUCLEOPHILIC SUBSTITUTION

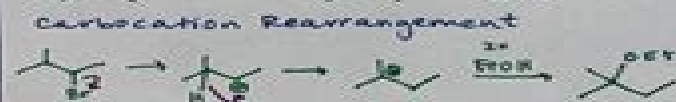
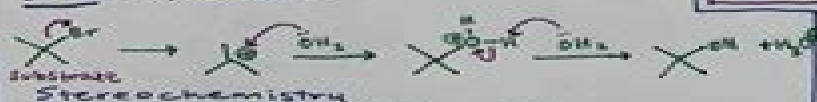


Nuc.
 Lone pairs
 (b) charge
 (b) or δ^+ charge

LEAVING GROUP: $\text{I} > \text{Br} > \text{Cl} \gg \text{F or OH}$

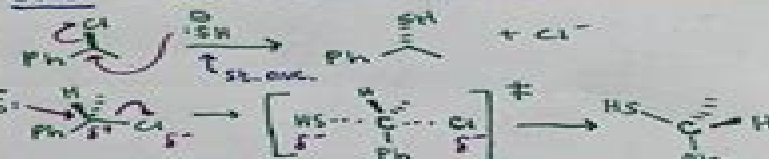


SN1 → Racemic



H or methyl shift, resonance
 * Weak Nucleophile, Rate = $k[\text{subst.}]$

SN2 → Invert chiral center



* Strong Nucleophile
 1° or methyl
 rate = $k[\text{substrate}][\text{nucleophile}]$

Guide Quick Notes Reactions Iitjee Organic Chemistry

Christian G. Meyer



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Table of Contents Guide Quick Notes Reactions Iitjee Organic Chemistry

1. Understanding the eBook Guide Quick Notes Reactions Iitjee Organic Chemistry
 - The Rise of Digital Reading Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guide Quick Notes Reactions Iitjee Organic Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Personalized Recommendations
 - Guide Quick Notes Reactions Iitjee Organic Chemistry User Reviews and Ratings
 - Guide Quick Notes Reactions Iitjee Organic Chemistry and Bestseller Lists
5. Accessing Guide Quick Notes Reactions Iitjee Organic Chemistry Free and Paid eBooks
 - Guide Quick Notes Reactions Iitjee Organic Chemistry Public Domain eBooks
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 - Guide Quick Notes Reactions Iitjee Organic Chemistry Budget-Friendly Options
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 - Guide Quick Notes Reactions Iitjee Organic Chemistry Enhanced eBook Features
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 - Highlighting and Note-Taking Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Interactive Elements Guide Quick Notes Reactions Iitjee Organic Chemistry
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guide Quick Notes Reactions Iitjee Organic Chemistry
 9. Balancing eBooks and Physical Books Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guide Quick Notes Reactions Iitjee Organic Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Setting Reading Goals Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Fact-Checking eBook Content of Guide Quick Notes Reactions Iitjee Organic Chemistry
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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