

Heat of Neutralization

Introduction:

The experiment was used to determine the heat evolved in an aqueous acid-base reaction. For strong acids and bases reacting in sufficiently dilute solutions, the heat evolved is approximately 57 kJ mol^{-1} at 298 K . Increasing the concentration should increase the heat of the reaction, while using a weaker acid or base will reduce the heat of the reaction. A series of acid-base neutralizations were performed on aqueous sodium hydroxide and the heat of the reaction was determined calorimetrically.

Experimental:

To measure the heat capacity of the calorimeter system, 225 mL of distilled water was added to a clean, dry Dewar. The water was stirred and the temperature read every 20 seconds until it became constant. Then, after recording the temperature of it, 25 mL of ice water was added to the flask. The mixture was stirred and the temperature was read until constant readings were obtained. The Dewar flask was then dried and 200 mL of the 0.2 N NaOH solution was added to the calorimeter. The temperature was taken before 50 mL of the 0.8 N HCl solution was added. The mixture was stirred and the temperature was read until constant readings were obtained. This process was then repeated twice for both the sulfuric and phosphoric acid solutions.

Results:

Calorimeter				Heat Capacity of Calorimeter	
Q	T _i	T _f	ΔT	$Q = mc\Delta T$	$C = \frac{Q}{\Delta T}$
$+102.5$	20.5°C	18.5°C	-2	$100 \times 4.18 \times (-2)$	$C = \frac{-836}{-2} = 418 \text{ J/K}$
$+200.3$	20.5°C	18.5°C	-2	$200 \times 4.18 \times (-2)$	$C = \frac{-1672}{-2} = 836 \text{ J/K}$
$+203.5$	20.5°C	18.5°C	-2	$200 \times 4.18 \times (-2)$	$C = \frac{-1672}{-2} = 836 \text{ J/K}$
$C_{\text{avg}} = 834 \text{ J/K}$					
HCl (0.8 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$Q = n\Delta H$	$\Delta H = \frac{Q}{n}$
$+102.5$	20.5°C	24.5°C	$+4$	$0.1 \times 57 \text{ kJ}$	$24.5 - 20.5 = 4$
$+200.3$	20.5°C	24.5°C	$+4$	$0.2 \times 57 \text{ kJ}$	$24.5 - 20.5 = 4$
$+203.5$	20.5°C	24.5°C	$+4$	$0.2 \times 57 \text{ kJ}$	$24.5 - 20.5 = 4$
H ₂ SO ₄ (0.4 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$Q = n\Delta H$	$\Delta H = \frac{Q}{n}$
$+102.5$	20.5°C	22.5°C	$+2$	$0.1 \times 57 \text{ kJ}$	$22.5 - 20.5 = 2$
$+200.3$	20.5°C	22.5°C	$+2$	$0.2 \times 57 \text{ kJ}$	$22.5 - 20.5 = 2$
$+203.5$	20.5°C	22.5°C	$+2$	$0.2 \times 57 \text{ kJ}$	$22.5 - 20.5 = 2$
H ₃ PO ₄ (0.3 N - 0.1 M)				Heat of Neutralization	
Q	T _i	T _f	ΔT	$Q = n\Delta H$	$\Delta H = \frac{Q}{n}$
$+102.5$	20.5°C	21.5°C	$+1$	$0.1 \times 57 \text{ kJ}$	$21.5 - 20.5 = 1$
$+200.3$	20.5°C	21.5°C	$+1$	$0.2 \times 57 \text{ kJ}$	$21.5 - 20.5 = 1$
$+203.5$	20.5°C	21.5°C	$+1$	$0.2 \times 57 \text{ kJ}$	$21.5 - 20.5 = 1$

Picture 1: Data tables and calculations for change in temperature and heat of neutralization

Heat Of Neutralization Lab Answer Key

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