

# Lab Report: Gravimetric Analysis of an Unknown Sulfate

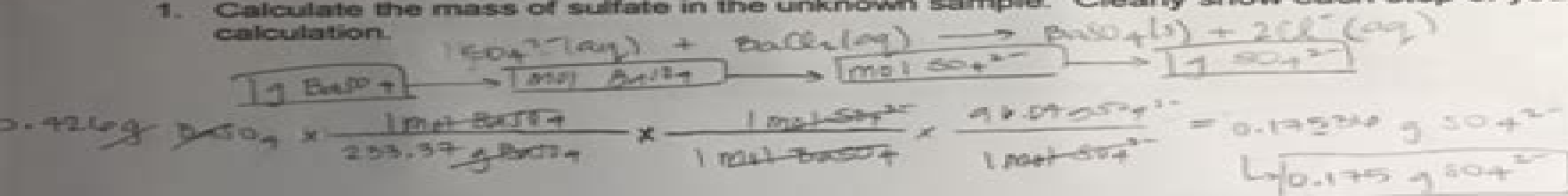
## Experimental Data

Unknown Sulfate ID Code: 20

|  |           |
|--|-----------|
| 1. Mass of empty 250-mL beaker                       | 111.244 g |
| 2. Mass of 250-mL beaker and unknown sulfate         | 112.227 g |
| 3. Mass of unknown sulfate                           | 0.343 g   |
| 4. Mass of empty crucible (without lid)              | 11.203 g  |
| 5. Mass of crucible (without lid) and barium sulfate | 11.629 g  |
| 6. Mass of barium sulfate                            | 0.426 g   |

## Calculations and Conclusions

- Calculate the mass of sulfate in the unknown sample. Clearly show each step of your calculation.



- Calculate the percentage by mass of sulfate in the unknown sample. Show your work.  
mass of unknown sulfate: 0.343 g ; mass of sulfate: 0.175 g  $\text{SO}_4^{2-}$

$$\% \text{SO}_4^{2-} = \frac{0.175 \text{ g}}{0.343 \text{ g}} \times 100 = 51.0204\%$$

$$\hookrightarrow \approx 51.0\%$$

- Calculate the percentage by mass of the metal in the unknown sample. Show your work.

$$100.0\% - 51.0\% = 49.0\%$$

# Gravimetric Analysis Lab Report Ucsd

**S Ashworth**



## **Gravimetric Analysis Lab Report Ucsd:**

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