



Which of the following substances is a reaction intermediate?

- A) N_2O_2 B) HO C) H_2O D) N_2O E) N_2O_2 and N_2O

15. The value of ΔH° for the reaction below is -126 kJ . _____ kJ is released when 2.00 mol of NaOH is formed in the reaction?



- A) 252 B) 63 C) 3.9 D) 7.8 E) -126

16. Estimate the enthalpy change for the reaction below from the average bond energies given. There are two $\text{C}-\text{Cl}$ and two $\text{C}-\text{H}$ bonds in CH_2Cl_2 . Remember that energy is absorbed when bonds are broken and released when they are formed.



Average Bond Energies

$\text{C}-\text{H}$ $+413 \text{ kJ/mol}$ $\text{Cl}-\text{Cl}$ $+242 \text{ kJ/mol}$

$\text{H}-\text{Cl}$ $+432 \text{ kJ/mol}$ $\text{C}-\text{Cl}$ $+339 \text{ kJ/mol}$

- A) -232 kJ/mol B) $+578 \text{ kJ/mol}$ C) $+232 \text{ kJ/mol}$
D) -578 kJ/mol E) $+541 \text{ kJ/mol}$

17. According to the reaction profile below, what is ΔH for the reaction: $2 \text{H}_2\text{O}(g) + 2 \text{Br}_2(g) \rightarrow 4 \text{HBr}(g) + \text{O}_2(g)$?



- A) -434 kJ B) $+434 \text{ kJ}$ C) -276 kJ D) $+276 \text{ kJ}$ E) $+158 \text{ kJ}$

18. Given the following at ΔH° (25°C and 1.00 atm):



Calculate the ΔH° for the reaction below at 25°C:



- A) $+11.0 \text{ kJ}$ B) $+44.3 \text{ kJ}$ C) $+55.3 \text{ kJ}$ D) -22.1 kJ E) -55.3 kJ

19. Consider the following PE diagram for a catalyzed and uncatalyzed reaction:

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Roman Wölfel



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