ACIDS, BASES AND SALTS

Definition of an acid	What causes the acidic properties of acids?		
An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	The hydrogen ions (H') cause the acidic properties and these are formed in the presence of water.		

Another term to refer to an acid	Why an acid is also called a proton donor?	
An acid is called a <u>proton donor</u>	It's because an acid provides protons or hydrogen ions (H1) to other	
	substances during the reaction.	

Substance to which an	A base is then called	Why the base is called a	Equation for the reaction between
acid provides protons		proton acceptor?	acid and base
BASE	PROTON ACCEPTOR	Because the base accepts hydrogen ions from acids	H'mg + OHmg H ₂ O _R

Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?	
Hydrochloric acid	HCI		They are derived from mineral salts ie chlorides for	
Sulphuric acid	H;80.	MINERAL ACIDS	HCl, sulphates for H ₂ SO ₄ and nitrates for HNO ₂	
Nitric acid	HNO ₃	STATES OF STATES		

Other minera known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H ₂ SO ₃	Derived from sulphites	Ethanoic acid	CITRIC ACID from lemons
Carbonic acid	H ₂ CO ₂	Derived from carbonates	(CH ₃ COOH)	TARTARIC ACID from grapes
Phosphoric acid	H ₂ PO ₄	Derived from phosphates		ACETIC ACID from vinegar
Nitrous acid	HNO ₃	Derived from nitrites	Methanoic acid	LACTIC ACID from sour milk
proceeding	Marine St.		(HCOOH)	Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces	Term given to the number of hydrogen ions produced by one molecule of an acid	Definition of basicity of an acid
HYDROGEN IONS	BASICITY OF AN ACID	BASICTTY of an acid is the <u>number of</u> <u>hydrogen ions</u> produced by <u>one</u> <u>molecule</u> of an acid <u>in aqueous</u> <u>solution</u> .

Guide Acids Bases And Salts

John Krenos, Joseph Potenza, Carl Hoeger

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows guick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Cambridge IGCSE® Chemistry Practical Teacher's** Guide with CD-ROM Michael Strachan, 2016-06-23 This edition of our successful series to support the Cambridge IGCSE Chemistry syllabus 0620 is fully updated for the revised syllabus from first examination from 2016 The Cambridge IGCSE Chemistry Practical Teacher's Guide complements the Practical Workbook helping teachers to include more practical work in lessons Specific support is provided for each of the carefully designed investigations to save teachers time The Teacher's Guide contains advice about planning investigations guidance about safety considerations differentiated learning suggestions to support students who might be struggling and to stretch the students who are most able as well as answers to all the questions in the Workbook The Teacher's Guide also includes a CD ROM containing model data to be used in instances when an investigation cannot be carried out Pocket Guide to Fluid, Electrolyte, and Acid-Base Balance - E-Book Ursula Heitz, Mima M. Horne, 2004-10-20 Compares different types of parenteral and enteral feedings along with nursing implications Contains expanded and updated relevant IV therapy and pharmacology Features updated content throughout with current literature and research findings such as most current blood pressure quidelines from the U S Department of Health and Human Services The Prospector's Field-book and Guide in the Search for and the Easy Determination of Ores and Other Useful Minerals Henry Stafford Osborn, 1901 Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is

certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications Practical/Laboratory Manual Chemistry Class XI based on NCERT guidelines by Dr. S. C. Rastogi & Er. Meera Goyal Dr. S. C. Rastogi , Er. Meera Goyal, 2020-06-23 An Excellent Book in Accordance with the latest syllabus for Class 11 Prescribed by CBSE NCERT and Adopted by Various State Education Boards A Basic Laboratory Techniques 1 To cut a glass tube or glass rod 2 To bend the glass rod at an angle 3 To draw a glass jet from a glass tube 4 To bore a cork and fit a glass tube into it B Characterisation and Purification of Chemical Substances 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method 3 To prepare crystals of pure potash alum K2SO4 Al2 SO4 3 24H2O from the given impure sample 4 To prepare the pure crystals of copper sulphate from the given crude sample 5 To prepare pure crystals of benzoic acid from a given impure sample C Measurement of pH Values 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper 2 To determine and compare the pH values of solutions of strong acid HCI and weak acid CH3COOH of same concentration 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper 4 To study the pH change by common ion CH3COO ion in case of weak acid CH3COOH 5 To determine the change in pH value of weak base NH4OH in presence of a common ion NH4 D Chemical Equilibrium 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions 2 To study the shift in equilibrium between Co H2O 6 2 and Cl ions by changing the concentrations of either of the ions E Quantitative Analysis 1 To prepare M 10 oxalic acid solution by direct weighing method 2 To prepare M 10 solution of sodium carbonate by direct weighing method 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution F Qualitative Analysis 1 Analysis of Anions 2 Analysis of Cations G Detection of

Elements in Organic Compounds 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne's test 2 To detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne's test INVESTIGATORY PROJECTS A Checking of Bacterial Contamination in Water 1 To check the bacterial contamination in drinking water by testing sulphide ions B Methods of Water Purification 1 To purify water from suspended impurities by using sedimentation 2 To purify water by boiling 3 To purify water by distillation method 4 To purify water by reverse osmosis technique 5 To purify water by GAC method 6 To purify water by bleach treatment 7 To purify water by oxidising agent 8 To purify water by ozone treatment method C Water Analysis 1 To test the hardness of different water samples D Foaming Capacity of Various Soaps 1 To compare the foaming capacity of different washing soaps 2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap E Tea Analysis 1 To study the acidity of different samples of tea leaves tea by using pH paper F Analysis of Fruits and Vegetable Juices 1 To analyse the fruit and vegetable juices for the constituent present in them G Rate of Evaporation 1 To study the rate of evaporation of different liquids H Effect of Acids and Bases on Tensile Strength of Fibres 1 To compare the tensile strength of natural fibres and synthetic fibres 2 To study the effect of acids and bases on tensile strength of different fibres Log Antilog Table Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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