



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In chemistry, it refers to a reaction where the forward and reverse rates are equal, and the concentrations of reactants and products remain constant over time. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the net change is zero.	Equilibrium Constant The equilibrium constant (K) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Graph showing concentration vs. time for a reaction at equilibrium. The concentration of reactants decreases and the concentration of products increases until they reach a constant value, indicating equilibrium.
Equilibrium Constant, K_c K_c is the equilibrium constant expressed in terms of molar concentrations. It is used for reactions in solution or in the gas phase where concentrations are known. For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_c is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p K_p is the equilibrium constant expressed in terms of partial pressures. It is used for reactions in the gas phase where partial pressures are known. For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_p is given by: $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
Equilibrium Constant, K K is the general equilibrium constant, which can be expressed in terms of either concentrations (K_c) or partial pressures (K_p). The value of K is constant for a given reaction at a specific temperature. For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Relationship between K_c and K_p The relationship between K_c and K_p is given by the equation: $K_p = K_c (RT)^{\Delta n}$ where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.	
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration	Effect on Equilibrium - Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the products. - Increasing the concentration of a product shifts the equilibrium to the left, favoring the reactants.	Result For a reaction where $\Delta n > 0$, increasing the concentration of a reactant increases the value of K_p. For a reaction where $\Delta n < 0$, increasing the concentration of a reactant decreases the value of K_p.
Pressure	Effect on Equilibrium - Increasing the pressure (by decreasing the volume) shifts the equilibrium towards the side with fewer moles of gas. - Decreasing the pressure (by increasing the volume) shifts the equilibrium towards the side with more moles of gas.	Result For a reaction where $\Delta n > 0$, increasing the pressure decreases the value of K_p. For a reaction where $\Delta n < 0$, increasing the pressure increases the value of K_p.
Temperature	Effect on Equilibrium - For an exothermic reaction, increasing the temperature shifts the equilibrium to the left, favoring the reactants. - For an endothermic reaction, increasing the temperature shifts the equilibrium to the right, favoring the products.	Result For an exothermic reaction, increasing the temperature decreases the value of K. For an endothermic reaction, increasing the temperature increases the value of K.

Guide Chemical Equilibrium

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Guide Chemical Equilibrium:

Embracing the Melody of Phrase: An Psychological Symphony within **Guide Chemical Equilibrium**

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