



At the end of this chapter, you will be able to describe the properties and uses of metals. You will see that the properties of metals are mainly determined by their structure. You will also understand how these properties are used to make products for many different purposes.

You will learn how metals are used to make products and to make structures in the industry of metals and how they are used to make products in the industry of metals and how they are used to make products in the industry of metals.

The properties of metals can be determined in the form of their reactions with acids, water and oxygen, and these reactions can be used to determine their properties.

Finally, you will learn how and why metals are used to make products and structures in the industry of metals and how they are used to make products in the industry of metals and how they are used to make products in the industry of metals.

Key knowledge

- The common properties of metals include: malleability, ductility, melting point, heat conductivity and electrical conductivity with reference to the nature of metallic bonding and the structure of metals (pages 4.1, 4.2).
- The common properties of a metal include: malleability, ductility, melting point, heat conductivity and electrical conductivity with reference to the nature of metallic bonding and the structure of metals (pages 4.1, 4.2).
- The common properties of a metal include: malleability, ductility, melting point, heat conductivity and electrical conductivity with reference to the nature of metallic bonding and the structure of metals (pages 4.1, 4.2).

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4.1 Metallic properties and bonding

More than 90% of the elements in the periodic table are metals (Figure 4.1.1).

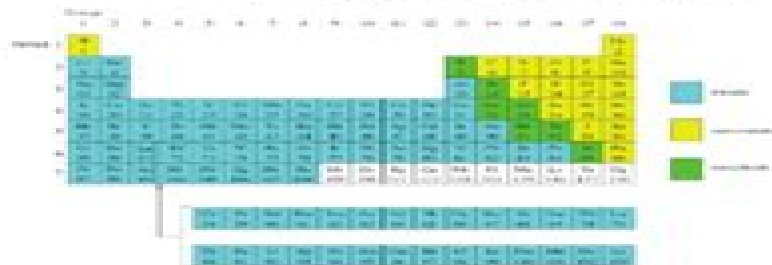


Figure 4.1.1 The distribution of metals, non-metals, and metalloids in the periodic table.

The development of metals in the periodic table is shown in Figure 4.1.1. The development of metals in the periodic table is shown in Figure 4.1.1. The development of metals in the periodic table is shown in Figure 4.1.1. The development of metals in the periodic table is shown in Figure 4.1.1.

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The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure.

Table 4.1.1 Properties and uses of some metals

Metals	Properties	Uses
Aluminium	Light, malleable, strong, good conductor of heat and electricity, resistant to corrosion, easily formed into sheets, pipes, rods, etc.	Construction, transport, packaging, electrical wiring, etc.
Iron	Strong, hard, malleable, good conductor of heat and electricity, resistant to corrosion, easily formed into sheets, pipes, rods, etc.	Construction, transport, packaging, electrical wiring, etc.
Steel	Strong, hard, malleable, good conductor of heat and electricity, resistant to corrosion, easily formed into sheets, pipes, rods, etc.	Construction, transport, packaging, electrical wiring, etc.

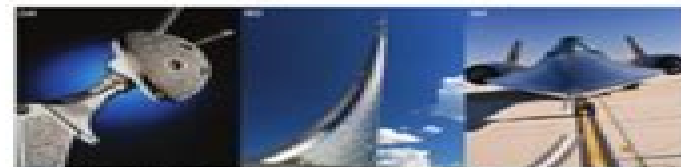


Figure 4.1.2 The use of metals in the construction of a large ship.

The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure. The properties of metals are determined by their structure.

GENERAL PROPERTIES OF METALS

Metals are characterized by their properties. Metals are characterized by their properties. Metals are characterized by their properties. Metals are characterized by their properties. Metals are characterized by their properties.

- Metals are good conductors of heat and electricity.
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