

LESSON

16

CLASSWORK

Old Gold

Formation of Elements

Name _____

Date _____ Period _____



Purpose

To practice interpreting nuclear equations and to predict the products of specific nuclear reactions.

Part I: Nuclear Equations

Use a periodic table and the isotope chart from Lesson 14.

1. Fill in this table.

Nuclear equation	Change in atomic number	Change in number of protons	Change in number of neutrons	Change in mass number	Type of nuclear change
a. ${}^{226}_{88}\text{Ra} \longrightarrow {}^4_2\text{He} + {}^{222}_{86}\text{Rn}$					alpha decay
b. ${}^{214}_{84}\text{Po} \longrightarrow {}^4_2\text{He} + {}^{210}_{82}\text{Pb}$					
c. ${}^{12}_6\text{C} + {}^{14}_7\text{N} \longrightarrow {}^{25}_{13}\text{P}$					
d. ${}^{40}_{20}\text{Ca} \longrightarrow {}^0_{-1}\text{e}^- + {}^{40}_{21}\text{Sc}$					
e. ${}^{152}_{64}\text{Gd} \longrightarrow {}^4_2\text{He} + {}^{148}_{62}\text{Sm}$				-4	
f. ${}^9_4\text{Be} + {}^4_2\text{He} \longrightarrow {}^{12}_6\text{C}$		+2			
g. ${}^{12}_6\text{C} \longrightarrow {}^0_{-1}\text{e}^- + {}^{12}_7\text{N}$	+1				beta decay

2. Based on the equations, how do you know which nuclear changes are fusion?

3. When an alpha particle is lost from an atom, where on the periodic table would you find the product of the nuclear reaction?

4. Explain how you figured out how many neutrons an isotope has.

5. Fill in the missing parts of these nuclear reactions:



Living By Chemistry Lesson 8 Unit 5

Clemens Wendtner



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