

Final Exam Study Guide (General Chemistry I)

- Chemistry - the study of matter and the changes that matter undergoes
- Matter - anything that has mass and takes up space
 - Three states of matter: solid, liquid, and gas (plasma is a gas)

	Proximity of Particles	Energetics	Size/Shape
Solid	Closely packed and in direct contact with each other	Have very little energy of motion (slow)	Fixed shape and fixed volume
Liquid	In indirect contact with each other; usually about a particles distance apart (bounce off each other)	Have an intermediate amount energy of motion (medium)	Does not have a fixed shape; has a fixed volume
Gas	Particles are very far apart	Have a large amount of energy of motion (fast)	No fixed shape or volume

- **PARTICLES ARE ALWAYS MOVING!**
- "Can you undo it?" (Page 12)
 - Physical Changes - changes physical appearance but not its composition
 - Ex. Melting, evaporation, freezing, etc.
 - Chemical Changes - changes in chemical composition of matter (chemical reaction)
 - Ex. Baking soda and vinegar, burning, mixing, etc.
 - Physical Properties - characteristics that can be described by observing the matter; no change in composition is required to detect a physical property
 - Ex. Color, odor, density, etc.
 - Chemical Properties - a change in composition is required to detect a chemical property
 - Ex. Flammability
- Accuracy Versus Precision:
 - In every measurement, there is some degree of uncertainty.
 - **NUMBERS WITHOUT UNITS ARE MEANINGLESS IN MEASUREMENT!**
 - Accuracy - the agreement of a measurement with a known or accepted value (correctness)
 - Precision - the agreement of a measurement with other measurements in a given set of measurements (reproducibility/repeatability)
 - You can have accuracy without precision, and precision without accuracy.
 - **WE ARE ONLY AS CERTAIN AS OUR LEAST CERTAIN NUMBER!**
 - 254 can mean 253.5 to 254.4
 - 254.0 can mean 253.95 to 254.04
 - 254.00 can mean 253.950 to 254.004
 - 254.000 can mean 253.9500 to 254.0004
- Density - the mathematical relationship of mass to volume

Grade 12 Chemistry Final Exam Study Guide

L Darling-Hammond



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Table of Contents Grade 12 Chemistry Final Exam Study Guide

1. Understanding the eBook Grade 12 Chemistry Final Exam Study Guide
 - The Rise of Digital Reading Grade 12 Chemistry Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 12 Chemistry Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 12 Chemistry Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 12 Chemistry Final Exam Study Guide
 - Personalized Recommendations
 - Grade 12 Chemistry Final Exam Study Guide User Reviews and Ratings
 - Grade 12 Chemistry Final Exam Study Guide and Bestseller Lists
5. Accessing Grade 12 Chemistry Final Exam Study Guide Free and Paid eBooks
 - Grade 12 Chemistry Final Exam Study Guide Public Domain eBooks
 - Grade 12 Chemistry Final Exam Study Guide eBook Subscription Services
 - Grade 12 Chemistry Final Exam Study Guide Budget-Friendly Options

6. Navigating Grade 12 Chemistry Final Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 12 Chemistry Final Exam Study Guide Compatibility with Devices
 - Grade 12 Chemistry Final Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 12 Chemistry Final Exam Study Guide
 - Highlighting and Note-Taking Grade 12 Chemistry Final Exam Study Guide
 - Interactive Elements Grade 12 Chemistry Final Exam Study Guide
8. Staying Engaged with Grade 12 Chemistry Final Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 12 Chemistry Final Exam Study Guide
9. Balancing eBooks and Physical Books Grade 12 Chemistry Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 12 Chemistry Final Exam Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 12 Chemistry Final Exam Study Guide
 - Setting Reading Goals Grade 12 Chemistry Final Exam Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 12 Chemistry Final Exam Study Guide
 - Fact-Checking eBook Content of Grade 12 Chemistry Final Exam Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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