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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Chemical reactions are represented by the reaction: $\rightleftharpoons$

What is Equilibrium? Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time having reached Equilibrium. Chemical reactions are represented by the reaction: $\rightleftharpoons$ Equilibrium Constant K_c Equilibrium Constant K_p	Equilibrium Constant Equilibrium Constant K_c Equilibrium Constant K_p	
Equilibrium Constant K_c Equilibrium Constant K_c is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. Equilibrium Constant K_c	Equilibrium Constant K_p Equilibrium Constant K_p is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient. Equilibrium Constant K_p	
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Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. Decreasing the concentration of a reactant will shift the equilibrium to the left, favoring the reactants.	Pressure Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.	Temperature Increasing the temperature will shift the equilibrium to the side that is endothermic. Decreasing the temperature will shift the equilibrium to the side that is exothermic.

Guide Chemical Equilibrium

Rosina Ehmann



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