

Thermochemistry Worksheet Answers

Useful Values:

Heat of Fusion for Water: $\Delta H_{\text{fus}} = 6 \text{ kJ/mol}$

Common error: not converting to moles

Heat of Vaporization for Water: $\Delta H_{\text{vap}} = 45 \text{ kJ/mol}$

Common error: same

1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K
Water (ice)	2.05	0.490
Water (liquid)	4.18	1.00
Water (gas)	2.08	0.498

Common error here:

using moles instead of grams.

1. How much energy does it take to warm the following substances the listed amount?

a. 20g of ice from -45°C to -15°C ?

294 calories / 1230 Joules

b. 20g of water from 25°C to 55°C ?

600 calories / 2508 Joules

c. 20g of water vapor from 130°C to 160°C ?

299 calories / 1249 Joules

2. To melt ice, one must bring ice to the freezing point (0°C), and then thaw it into water. How much heat would it take to melt...

a. 30g of ice at 0°C ?

10 kJ

(10,000 J)

b. 30g of ice at -40°C ?

12.46 kJ

(12,460 J)

3. To boil water, one must raise it to the boiling point (100°C), and then boil it into steam. How much heat would it take to boil...

a. 40g of water at 100°C ?

100 kJ

(100,000 Joules)

b. 40g of water at 90°C ?

101.7 kJ

(101,672 Joules)

c. 40g of water at 0°C ?

116.7 kJ

(116,720 Joules)

d. 40g of ice at 0°C ?

130.1 kJ

(130,053 Joules)

e. 40g of ice at -80°C ?

136.6 kJ

(136,613 Joules)

4. If 700 calories of heat raise 400g of Phonium by 20°C , what is the specific heat of Phonium? **0.0875 cal/g/K**

5. How much heat energy is stored in 2g of ice at -30°C ? (remember: use Kelvin)

238 cal / 996 J

6. How much heat energy is stored in 2g of Phonium at 70°C ? (Assume Phonium is a solid at 70°C) **1672 cal**

7. If 100g of water at 10°C comes in contact with _____, what will the final temperature be?

(Remember, you have the specific heat of Phonium from question 4.)

a. 100g of water at 90°C

50°C

b. 100g of Phonium at 90°C (Hint: the water and Phonium end at the same temperature) **16.44°C**

Common error: 6.44 degrees is the temperature change, and the water gets warmer.

c. 20g of Phonium at 90°C

11.38°C

d. 1g of Phonium at 90°C

10.07°C

8. If 450g of Phakelite at 100°C is added to 1000g of water at 20°C , the temperature of the water increases by 5° . What is the specific heat of Phakelite? **0.148 cal/g/K**

8. If 450g of Nonsensium at 500°C is added to 1000g of water at 20°C , the temperature of the water increases by 25° . What is the specific heat of Nonsensium? **0.122 cal/g/K**

Guided And Study Workbook Thermochemistry Answers

**Barron's Educational Series, Mark
Kernion, Joseph A. Mascetta**



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