

HL Paper 2

Enthalpy changes depend on the number and type of bonds broken and formed.

The table lists the standard enthalpies of formation, $\Delta_f H^\ominus$, for some of the species in the reaction above.

	$\text{CH}_4(\text{g})$	$\text{H}_2\text{O}(\text{g})$	$\text{CO}(\text{g})$	$\text{H}_2(\text{g})$
$\Delta_f H^\ominus / \text{kJ mol}^{-1}$	-74.0	-242	-111	

Enthalpy changes depend on the number and type of bonds broken and formed.

a. Hydrogen gas can be formed industrially by the reaction of natural gas with steam. [3]



Determine the enthalpy change, ΔH , for the reaction, in kJ, using section 11 of the data booklet.

Bond enthalpy for $\text{C}\equiv\text{O}$: 1077 kJ mol^{-1}

b.i. Outline why no value is listed for $\text{H}_2(\text{g})$. [1]

b.ii. Determine the value of $\Delta H^\ominus$, in kJ, for the reaction using the values in the table. [1]

c. The table lists standard entropy, $S^\ominus$, values. [1]

	$\text{CH}_4(\text{g})$	$\text{H}_2\text{O}(\text{g})$	$\text{CO}(\text{g})$	$\text{H}_2(\text{g})$
$S^\ominus / \text{J K}^{-1} \text{mol}^{-1}$	+186	+189	+108	+131

Calculate the standard entropy change for the reaction, $\Delta S^\ominus$, in J K^{-1} .



d. Calculate the standard free energy change, $\Delta G^\ominus$, in kJ, for the reaction at 298 K using your answer to (b)(ii). [1]

e. Determine the temperature, in K, above which the reaction becomes spontaneous. [1]

Markscheme

a. bonds broken: $4(\text{C}-\text{H}) + 2(\text{H}-\text{O})/4(414) + 2(463)/2582 \text{ kJ}$

bonds made: $3(\text{H}-\text{H}) + \text{C}\equiv\text{O}/3(436) + 1077/2385 \text{ kJ}$

$$\Delta H = \text{IBE}_{\text{bonds broken}} - \text{IBE}_{\text{bonds made}} = 2582 - 2385 = -197 \text{ kJ}$$

Award [3] for correct final answer.

Award [2 max] for -197 kJ.

[3 marks]

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Mark Steyvers



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