CHAPTER 6

1. The energy eigenvalues are $\epsilon_k = \frac{h^2}{2m} k^2$. The mean value over the volume of a sphere in k space is

$$<\epsilon> = \frac{\dot{h}^2}{2m} \frac{\int k^2 dk \cdot k^2}{\int k^2 dk} = \frac{3}{5} \cdot \frac{\ddot{h}^2}{2m} k_{_F}{}^2 = \frac{3}{5} \epsilon_{_F}. \label{epsilon}$$

The total energy of N electrons is

$$U_n = N \cdot \frac{3}{5} \epsilon_F.$$

In general p = -∂U/∂V at constant entropy. At absolute zero all processes are at constant entropy (the

Third Law), so that $p = -dU_e/dV$, where $U_e = \frac{3}{5}N\epsilon_F = \frac{3}{5}N\frac{\hbar^2}{2m}\left(\frac{3\pi^2N}{V}\right)^{2/3}$, whence

 $p = \frac{2}{3} \cdot \frac{U_0}{V}$, (b) Bulk modulus

$$B = -V \frac{dp}{dV} = V \left(-\frac{2}{3} \frac{U_0}{V^2} + \frac{2}{3V} \frac{dU_0}{dV} \right) = \frac{2}{3} \cdot \frac{U_0}{V} + \left(\frac{2}{3} \right)^2 \frac{U_0}{V} = \frac{10}{9} \frac{U_0}{V}.$$

$$\frac{U_0}{V} = \frac{3}{5} (4.7 \times 10^{22} \text{ cm}^{-3}) (4.7 \text{ eV}) (1.6 \times 10^{-12} \text{ erg/eV})$$
$$= 2.1 \times 10^{11} \text{ erg cm}^{-3} = 2.1 \times 10^{11} \text{ dyne cm}^{-2},$$

whence $B = 2.3 \times 10^{11}$ dyne cm⁻². By experiment (Table 3.3), $B = 1.2 \times 10^{11}$ dyne cm⁻².

3. The number of electrons is, per unit volume, $n = \int_0^{\infty} d\epsilon \ D(\epsilon) \cdot \frac{1}{e^{(\epsilon - \mu)/\epsilon} + 1}$, where $D(\epsilon)$ is the density of orbitals. In two dimensions

$$n = \frac{m}{\pi \hbar^2} \int_0^{\infty} d\epsilon \frac{1}{e^{i\epsilon - \mu h/\tau} + 1}$$

$$= \frac{m}{\pi \hbar^2} (\mu + \tau \log (1 + e^{-\mu/\tau})),$$

where the definite integral is evaluated with the help of Dwight [569.1].

4a. In the sun there are $\frac{2 \times 10^{33}}{1.7 \times 10^{-24}} = 10^{42}$ mucleons, and roughly an equal number of electrons. In a white dwarf star of volume

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Padhraic Smyth

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Molecular Physics and Elements of Quantum Chemistry Hermann Haken, Hans Christoph Wolf, 2013-03-09 This textbook is intended for use by students of physics physical chemistry and theoretical chemistry The reader is presumed to

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