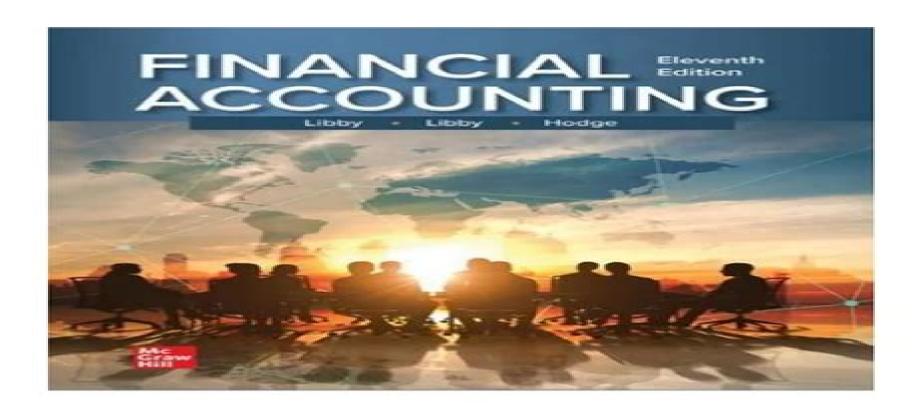
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water n a 54 g 18 g mol 1 3 mol total number of moles n a n b 0 5 3 3 5 mol mole fraction of solute ethyl alcohol x b n b n a n b 0 5 3 5 0 1429

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g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass of solvent water 100 g 0 1 kg to find molarity of solution

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction mltiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature