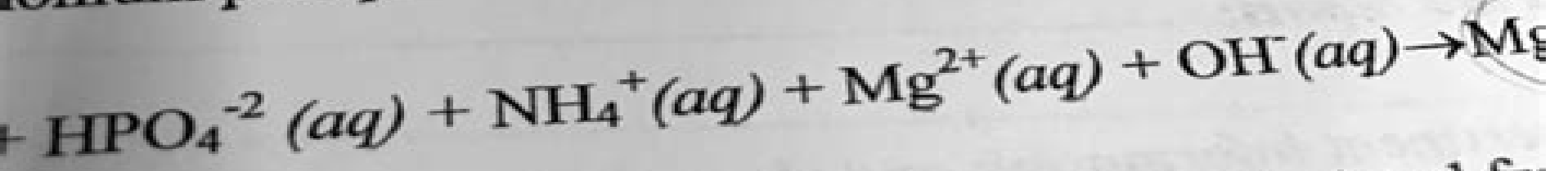


## Experiment for Gravimetric Analysis:

### The Determination of Phosphorus in Plant Food

from "Laboratory Experiments", J.H. Nelson and K.C. Kemp, Prentice-

Gravimetric analysis is a quantitative (i.e. how much?) method of classical analysis. The element to be determined is isolated in a solid compound of known identity. The mass of the element that was present in the original sample can be calculated from the mass of this compound. Plant foods contain three essential nutrients that are soluble in water. These are soluble compounds of nitrogen, phosphorus, and potassium. A plant food will have a set of numbers such as 15-30-15. This means that the product is guaranteed to contain at least 15% nitrogen, 30% phosphorus, and 15% potassium (expressed as  $K_2O$ ). The remaining of the product is made up of anions and cations to balance the charge in the chemical compound. This experiment is a quality control analysis for the determination of phosphorus in a plant food. Phosphorus will be determined by precipitation of ammonium phosphate hexahydrate according to the reaction:



The percentage of  $P_2O_5$  in the initial sample can be calculated from the mass of the precipitate.

# Gravimetric Determination Of Phosphorus Lab

**Michael Brown**



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