

- How many moles of **atoms** are there in:
 - 64 g of copper
 - 28 g of iron
- How many moles of **molecules** are there in:
 - 4.5 g of H_2O
 - 1.9 g of fluorine gas.
- What is the mass of:
 - 3 moles of silver
 - 0.01 mole of V_2O_5
- How many moles of **ions** are present in:
 - 3 mole of KBr
 - 0.02 moles of MgBr_2
- How many moles of **atoms** are present in:
 - 0.4 moles of chlorine gas
 - 1 mole of carbon dioxide
- What is the empirical formula for a compound with molecular formula $\text{C}_6\text{H}_8\text{N}_2\text{O}_4$?
- Copy and balance the following equation:

$$\text{.....As}_4 + \text{.....O}_2 \rightarrow \text{.....As}_4\text{O}_{10}$$
- A hydrocarbon has the empirical formula CH_2 and a molar mass of 70 g/mol. What is its molecular formula?
- A compound has the following percentage composition:
 K: 38.6 % N: 13.9 % O: 47.5 %
 Work out the empirical formula of the compound.
- Calculate the percentage by mass of each element in NH_4NO_3 .
- Below are the results from an experiment carried out to find the formula of an oxide of copper. In this experiment some copper metal was weighed before and after heating to constant mass in a crucible.

Mass of crucible	= 15.6 g
Mass of crucible + copper	= 18.8 g
Mass of crucible + copper oxide	= 19.2 g

 - What is the mass of copper used?
 - What is the mass of copper oxide produced?
 - What is the mass of oxygen that combined with copper?
 - Using your answers to parts (a) and (c), calculate the empirical formula of the copper oxide (a table will be useful).

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