

actomis

(a) is not important because nuclear forces are short-ranged

(b) is as important as electrostatic force for binding the two atoms

(c) cancels the repulsive electrostatic force between the nuclei

(d) is not important because oxygen nucleus have equal number of neutrons and protons

Answers

(a) is not important because nuclear forces are short-ranged

12.6. Two H atoms in the ground state collide inelastically. The maximum amount by which their combined kinetic energy is reduced is

(a) 10.20 eV

(b) 20.40 eV

(c) 13.6 eV

(d) 27.2 eV

Answer:

(a) 10.20 eV

12.7. A set of atoms in an excited state decays.

(a) in general to any of the states with lower energy

(b) into a lower state only when excited by an external electric field

(c) all together simultaneously into a lower state

(d) to emit photons only when they collide

Answer:

(a) in general to any of the states with lower energy

#### Multiple Choice Questions II

12.8. An ionised H-molecule consists of an electron and two protons. The protons are separated by a small distance of the order of angstrom. In the ground state,

(a) the electron would not move in circular orbits

(b) the energy would be (2)4 times that of a H-atom

(c) the electrons, orbit would go around the protons

(d) the molecule will soon decay in a proton and a H-atom

Answer:

(a) the electron would not move in circular orbits

(c) the electrons, orbit would go around the protons

12.9. Consider aiming a beam of free electrons towards free protons. When they scatter, an electron and a proton cannot combine to produce a H-atom,

(a) because of energy conservation

(b) without simultaneously releasing energy in the from of radiation

(c) because of momentum conservation

(d) because of angular momentum conservation Answer:

(a) because of energy conservation

(b) without simultaneously releasing energy in the from of radiation

12.10. The Bohr model for the spectra of a H-atom

(a) will not be applicable to hydrogen in the molecular from

(b) will not be applicable as it is for a He-atom

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**Clifford Lane** 

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- 1. Understanding the eBook Grade 12 Physics 2014 Exemplar
  - The Rise of Digital Reading Grade 12 Physics 2014 Exemplar
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Grade 12 Physics 2014 Exemplar
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade 12 Physics 2014 Exemplar
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Grade 12 Physics 2014 Exemplar
  - Personalized Recommendations
  - Grade 12 Physics 2014 Exemplar User Reviews and Ratings
  - Grade 12 Physics 2014 Exemplar and Bestseller Lists

- 5. Accessing Grade 12 Physics 2014 Exemplar Free and Paid eBooks
  - Grade 12 Physics 2014 Exemplar Public Domain eBooks
  - Grade 12 Physics 2014 Exemplar eBook Subscription Services
  - Grade 12 Physics 2014 Exemplar Budget-Friendly Options
- 6. Navigating Grade 12 Physics 2014 Exemplar eBook Formats
  - o ePub, PDF, MOBI, and More
  - Grade 12 Physics 2014 Exemplar Compatibility with Devices
  - Grade 12 Physics 2014 Exemplar Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Grade 12 Physics 2014 Exemplar
  - Highlighting and Note-Taking Grade 12 Physics 2014 Exemplar
  - Interactive Elements Grade 12 Physics 2014 Exemplar
- 8. Staying Engaged with Grade 12 Physics 2014 Exemplar
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade 12 Physics 2014 Exemplar
- 9. Balancing eBooks and Physical Books Grade 12 Physics 2014 Exemplar
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade 12 Physics 2014 Exemplar
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 12 Physics 2014 Exemplar
  - Setting Reading Goals Grade 12 Physics 2014 Exemplar
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 12 Physics 2014 Exemplar
  - Fact-Checking eBook Content of Grade 12 Physics 2014 Exemplar
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

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- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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