## Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
  - a. Write the balanced equation for the reaction given above:

 $CuCl_2 + NaNO_3 \rightarrow Cu(NO_3)_2 + NaCl$  $CuCl_2 + 2 NaNO_3 \rightarrow Cu(NO_3)_2 + 2 NaCl$ 

b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

15 g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 134.362g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaCl

20 g NaNO<sub>3</sub> x 1 mol NaNO<sub>3</sub> x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84.997g NaNO<sub>3</sub> 2 mol NaNO<sub>3</sub> 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (11) chloride
- d. How many grams of copper (II) nitrate is formed?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 1 mol Cu(NO<sub>3</sub>)<sub>2</sub> x 187.467 g Cu(NO<sub>3</sub>)<sub>2</sub> = 20.9 g Cu(NO<sub>3</sub>)<sub>2</sub> 134.452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol Cu(NO<sub>3</sub>)<sub>2</sub>

e. How much of the excess reagent is left over in this reaction?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84,997g NaNO<sub>3</sub> = 19 g NaNO<sub>3</sub> 134,452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaNO<sub>3</sub>

20 g NaNO<sub>3</sub> - 19 g NaNO<sub>3</sub> = 1 g NaNO<sub>3</sub>

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction? Percent yield = actual yield / theoretical yield × 100 percent.

11.3/13.0 x 100% = 86.9%

- When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.
  - a. Balance the following equation:

 $Pb(NO_3)_{2 \text{ (sq)}} + NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + NaNO_{3 \text{ (sq)}}$  $Pb(NO_3)_{2 \text{ (sq)}} + 2 NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + 2 NaNO_{3 \text{ (sq)}}$ 

b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO<sub>3</sub>)<sub>2</sub> x 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 12.8 g NaNO<sub>3</sub> 331.214 g Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol NaNO<sub>3</sub>

15.0 g Nal x 1 mol Nal x 2 mol NaNO<sub>5</sub> x 84.997g NaNO<sub>5</sub> = 8.51 g NaNO<sub>5</sub> 149.89g Nal 2 mol Nal 1 mol NaNO<sub>5</sub>

8.51 g NaNO3 can be made

# **Limiting Reagent And Percent Yield Answers**

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## **Limiting Reagent And Percent Yield Answers:**

Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard,2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading

Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation STOICHIOMETRY NARAYAN CHANGDER, 2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel https www youtube com smartquiziz I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today s academic environment

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