

Types of Chemical Reactions Worksheet #1

H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2 \text{MgO}$ **Synthesis**
- 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
- 3) $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$ **Double Replacement**
- 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
- *5) $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
- 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3 \text{O}_2$ **Decomposition**
- 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
- 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
- 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
- *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**

*Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2 \text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2 \text{NaBr} \rightarrow 2 \text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3 \text{O}_2$
- 14) Synthesis: $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
- 15) Double replacement: $3 \text{Ca(OH)}_2 + 2 \text{FeCl}_3 \rightarrow 2 \text{Fe(OH)}_3 + 3 \text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2 \text{Hg}_2\text{O} \rightarrow 4 \text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2 \text{AgNO}_3 \rightarrow 2 \text{Ag} + \text{Cu(NO}_3)_2$ [Copper (II) is used here]

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Table of Contents Guide 16 Solution Chemistry Answer Key

1. Understanding the eBook Guide 16 Solution Chemistry Answer Key
 - The Rise of Digital Reading Guide 16 Solution Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Guide 16 Solution Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guide 16 Solution Chemistry Answer Key
 - User-Friendly Interface
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 - Personalized Recommendations
 - Guide 16 Solution Chemistry Answer Key User Reviews and Ratings
 - Guide 16 Solution Chemistry Answer Key and Bestseller Lists
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 - Highlighting and Note-Taking Guide 16 Solution Chemistry Answer Key
 - Interactive Elements Guide 16 Solution Chemistry Answer Key
- 8. Staying Engaged with Guide 16 Solution Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guide 16 Solution Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Guide 16 Solution Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guide 16 Solution Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guide 16 Solution Chemistry Answer Key
 - Setting Reading Goals Guide 16 Solution Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guide 16 Solution Chemistry Answer Key
 - Fact-Checking eBook Content of Guide 16 Solution Chemistry Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web jun 1 2017 *threat of nuclear war as clichéd as it may sound the contestants can expect questions about the threat of war or the need for global peace but please avoid giving pageant patty world peace answers we suggest that contestants should make themselves aware about at least a few basic facts about nuclear war or threat from north*

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