

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_

**a.** ionic bond

\_\_\_\_\_ **bond energy** \_\_\_\_\_

**b.** covalent bond

\_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_

**c.** metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_

**a.** H—F

\_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_

**d.** H—H

\_\_\_\_\_ **ionic** \_\_\_\_\_

**b.** Na—Cl

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

**e.** H—C

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

**c.** H—O

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

**f.** H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Guide Covalent Bonding Answer Key

**Robert R. Redfield**



## **Guide Covalent Bonding Answer Key:**

## Reviewing **Guide Covalent Bonding Answer Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Guide Covalent Bonding Answer Key**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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### **Table of Contents Guide Covalent Bonding Answer Key**

1. Understanding the eBook Guide Covalent Bonding Answer Key
  - The Rise of Digital Reading Guide Covalent Bonding Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Guide Covalent Bonding Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guide Covalent Bonding Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guide Covalent Bonding Answer Key
  - Personalized Recommendations
  - Guide Covalent Bonding Answer Key User Reviews and Ratings
  - Guide Covalent Bonding Answer Key and Bestseller Lists

5. Accessing Guide Covalent Bonding Answer Key Free and Paid eBooks
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  - Guide Covalent Bonding Answer Key Budget-Friendly Options
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  - Guide Covalent Bonding Answer Key Compatibility with Devices
  - Guide Covalent Bonding Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guide Covalent Bonding Answer Key
  - Highlighting and Note-Taking Guide Covalent Bonding Answer Key
  - Interactive Elements Guide Covalent Bonding Answer Key
8. Staying Engaged with Guide Covalent Bonding Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guide Covalent Bonding Answer Key
9. Balancing eBooks and Physical Books Guide Covalent Bonding Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guide Covalent Bonding Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guide Covalent Bonding Answer Key
  - Setting Reading Goals Guide Covalent Bonding Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guide Covalent Bonding Answer Key
  - Fact-Checking eBook Content of Guide Covalent Bonding Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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