

General Chemistry II Jasperse
Chemical equilibria. Extra Practice Problems

General Types/Groups of problems:

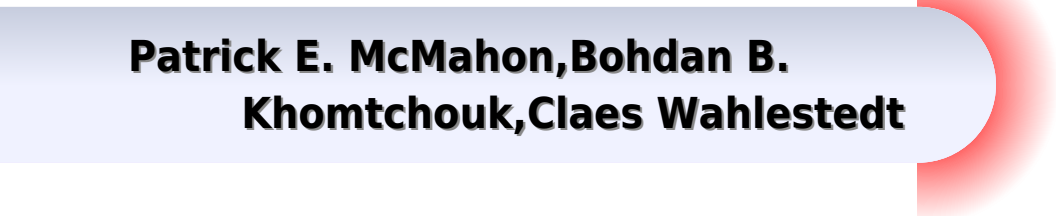
Equilibrium Conceptual	p1	Using Iee: Generic, Then Real But Simple Numbers	p8
Writing the Equilibrium Constant	p3	Solving for K given Initial and at Least one Equilibrium Concentration	p9
Manipulations of K: Reversing or Multiplying	p5	Solving for Equilibrium Concentrations Using Iee, Given K	p10
Solving for K Given All Equilibrium Concentrations	p6	LeChatelier's Principle, Q, and Responses to Disrupted Equilibria	p12
Solving for an Equilibrium Concentration Given K and Other Equilibrium Concentrations	p7	Answers	p15

Big-Picture Introductory Conceptual Questions

- Which of the following is true for a chemical reaction at equilibrium?
 - only the forward reaction stops
 - only the reverse reaction stops
 - both the forward and reverse reactions stop
 - the rate constants for the forward and reverse reactions are equal
 - the rates of the forward and reverse reactions are equal
- Which of the following is true regarding the **concentration of products**, for a chemical reaction **that is already at equilibrium**, assuming no disruptions to the equilibrium?
 - The concentrations of products will not change because there are no more reactants.
 - The concentrations of products will not change because the limiting reagent is gone.
 - The concentrations of products will not change because the forward and reverse rates are equal.
 - The concentrations of products will change continually because of reversibility.
- Which of the following are equal for a chemical system at equilibrium? If all are equal, answer E.
 - the **concentrations** of reactant and products are equal
 - the **rate constants** for the forward and reverse reactions are equal
 - the **time** that a particular atom or molecule spends as a reactant and product are equal
 - the **rate** of the forward and reverse reaction
 - all** of the above are equal
- A chemical equilibrium may be established by starting a reaction with _____.
 - reactants only.
 - products only.
 - equal quantities of reactants and products.
 - any quantities of reactants and products.
 - all the above
- An equilibrium that strongly favors products has _____.
 - a value of $K \ll 1$.
 - a value of $K \gg 1$.
 - a value of $Q \gg 1$.
 - a value of $Q \ll 1$.
 - $K = Q$.

Guided Practice Problems Chemistry Ch 16

**Patrick E. McMahon, Bohdan B.
Khomtchouk, Claes Wahlestedt**



Guided Practice Problems Chemistry Ch 16:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

Survival Guide to Organic Chemistry Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions

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