

## Final Exam Study Guide (General Chemistry I)

- Chemistry - the study of matter and the changes that matter undergoes
- Matter - anything that has mass and takes up space
  - Three states of matter: solid, liquid, and gas (plasma is a gas)

	Proximity of Particles	Energetics	Size/Shape
Solid	Closely packed and in direct contact with each other	Have very little energy of motion (slow)	Fixed shape and fixed volume
Liquid	In indirect contact with each other; usually about a particles distance apart (bounce off each other)	Have an intermediate amount energy of motion (medium)	Does not have a fixed shape; has a fixed volume
Gas	Particles are very far apart	Have a large amount of energy of motion (fast)	No fixed shape or volume

- **PARTICLES ARE ALWAYS MOVING!**
- "Can you undo it?" (Page 12)
  - Physical Changes - changes physical appearance but not its composition
    - Ex. Melting, evaporation, freezing, etc.
  - Chemical Changes - changes in chemical composition of matter (chemical reaction)
    - Ex. Baking soda and vinegar, burning, mixing, etc.
  - Physical Properties - characteristics that can be described by observing the matter; no change in composition is required to detect a physical property
    - Ex. Color, odor, density, etc.
  - Chemical Properties - a change in composition is required to detect a chemical property
    - Ex. Flammability
- Accuracy Versus Precision:
  - In every measurement, there is some degree of uncertainty.
  - **NUMBERS WITHOUT UNITS ARE MEANINGLESS IN MEASUREMENT!**
  - Accuracy - the agreement of a measurement with a known or accepted value (correctness)
  - Precision - the agreement of a measurement with other measurements in a given set of measurements (reproducibility/repeatability)
  - You can have accuracy without precision, and precision without accuracy.
  - **WE ARE ONLY AS CERTAIN AS OUR LEAST CERTAIN NUMBER!**
    - 254 can mean 253.5 to 254.4
    - 254.0 can mean 253.95 to 254.04
    - 254.00 can mean 253.950 to 254.004
    - 254.000 can mean 253.9500 to 254.0004
- Density - the mathematical relationship of mass to volume

# Grade 12 Chemistry Final Exam Study Guide

**Jiyuan Zhang**



## **Grade 12 Chemistry Final Exam Study Guide:**

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discipline This section is organized so as to parallel the order of topics found in most introductory psychology textbooks Intended for academicians who teach the introductory psychology course and or oversee grad assistants who teach the course all royalties of the book go directly to the Society for the Teaching of Psychology to promote its activities to further improve the teaching of psychology

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