

Appendix E

Math and Chemistry

In chemistry, quantities and values are often very, very large or very, very small. One mole of H_2O contains a huge number (602 214 199 000 000 000 000 000) of molecules. The mass of each molecule is minuscule (0.000 000 000 000 000 000 000 029 g). To calculate the mass of one mole of H_2O using the above data, you multiply the number of molecules in a mole by the mass of an individual molecule. Imagine doing this by hand:

[illegible]

There is another problem with writing all these zeros: It implies that they are all significant digits. Yet the non-final zeros to the right of the decimal point in the mass of an H_2O molecule are used as placeholders. They are not significant digits. There are only three significant digits in

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Scientific Notation

To simplify reporting large numbers and doing calculations, you can use scientific notation. As you learned in an earlier mathematics course, scientific notation works by using powers of 10 as multipliers.

The number of particles in one mole of a substance (the Avogadro constant) is often rounded down to 602 000 000 000 000 000 000 000. When performing calculations, however, this number is still too unwieldy. To simplify it, you can express it as a number between 1 and 10, multiplied by factors of 10. To do this, move the decimal point behind the left-most non-zero digit, counting the number of places that the decimal point moves. The number of places moved is the exponent for the base 10.

[illegible]

FIGURE E-1 The decimal point moves to the left.

Numbers less than 1 are converted into scientific notation by moving the decimal point to the right of the left-most non-zero digit. The exponent is expressed as a negative power of 10.

0.000 000 000 000 000 000 000 000 02.99 g

3 6 9 12 15 18 21

2.99×10^{-22}

Figure F.3 The decimal point moves to the right.

Once you have numbers in scientific notation, calculations become much simpler. To find the mass of one mole of H_2O , you can multiply $(6.02 \times 10^{23}/\text{mol}) \times (2.99 \times 10^{-23} \text{ g})$. To do this without a calculator, you first multiply:

$$6.02 \times 2.99 = 17.9998 \rightarrow 18.0 \text{ (rounded up to three significant digits)}$$

When multiplying, exponents are added to find the new exponent: $-22 + 21 = -1$.

This gives 10.0×10^6 g/mol. In scientific notation, however, there must be only one digit before the decimal place. Here the decimal place must be moved one place to the left: 1.00×10^7 g/mol.

Figure E.3 shows how to do this calculation with a scientific calculator. When you enter an exponent on a scientific calculator, you do not have to enter D 300 at any time.

Keystrokes	Display
$\frac{\square}{\square}$ \cdot \square \square \square \square \square \square	6.02E-23
\square \square \cdot \square \square \square \square \square \square	6.02E-23
\square	0.000

Round to three significant digits and express in scientific notation: 1.80×10^3 g/mol

Figure 6.3 On some scientific calculators, the $\boxed{\text{---}}$ key is labelled $\boxed{\text{---}}$. Key in negative exponents by entering the exponent, then striking the $\boxed{\text{---}}$ key.

Rules for Scientific Notation

- When multiplying, exponents are added abstracally

$$\begin{aligned} (2.37 \times 10^2 \text{ g}) &\times (1.89 \times 10^{-1} \text{ g}) \\ &= (2.37 \times 1.89) \times 10^{(2 + (-1))} \\ &= 4.4793 \times 10^1 \text{ g} \rightarrow 4.48 \times 10^1 \text{ g} \end{aligned}$$

- When dividing, exponents are subtracted algebraically.

$$\begin{aligned} (9.23 \times 10^2 \text{ L}) + (4.11 \times 10^3 \text{ L}) \\ = (9.23 + 4.11) \times 10^3 \text{ L} \\ = 13.345742 \times 10^3 \text{ L} \rightarrow 1.33 \times 10^4 \text{ L} \end{aligned}$$

- When adding or subtracting numbers in scientific notation, the numbers must be converted to the same power of 10 as the measurement with the greatest power of 10. Once the numbers are all expressed to the same power of 10, the power of 10 is neither added nor subtracted in the calculation.

$$\begin{aligned} & (8.22 \times 10^3 \text{ cm}) + (5.66 \times 10^3 \text{ cm}) \\ &= (8.22 \times 10^3 \text{ cm}) + (5.66 \times 10^3 \text{ cm}) \\ &\approx 8.786 \times 10^3 \text{ cm} \rightarrow 8.79 \times 10^3 \text{ cm} \end{aligned}$$

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