

Name _____

Chemistry Period _____

Worksheet Specific Heat

Specific Heat of water = $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from 25.5°C to 52.5°C . In the process, 3050J of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is $0.140 \text{ J/g}^\circ\text{C}$. **ANS: 807 g**
2. A block of aluminum with a mass of 140g is cooled from 98.4°C to 62.2°C with a release of 1137J of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of 226 J of heat are absorbed as 58.3 g of lead is heated from 12.0°C to 42.0°C . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when 132g of copper at 32.2°C is raised to 45.0°C ? The specific heat of copper is $0.385 \text{ J/g}^\circ\text{C}$. **ANS: 650 J**
5. What is the change in heat energy when 40.0g of water at 25.0°C is raised to 60.0°C ? **ANS: 5852 J**
6. What is the change in heat energy when 75.0 of water at 9.8°C is raised to 22.4°C ? **3950 J**
7. What is the change in heat energy when 125g of water at 10.0°C is raised to 90.0°C ? **ANS: 41.8 kJ**
8. What is the change in heat energy when 64.82 g of aluminum metal at 100.0°C is cooled to 82.0°C ? The specific heat of aluminum is $0.897 \text{ J/g}^\circ\text{C}$. **ANS: 1047**
9. What is the change in temperature if 24g of water at 50°C absorbs 985 J of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost 2300J of heat energy when it cooled from 80°C to 30°C ? The specific heat of iron is $0.449 \text{ J/g}^\circ\text{C}$. **ANS: 102 g**
11. What is the specific heat of metal if its mass is 26.86g and it requires 418.6 J of heat energy to raise its temperature from 27.4°C to 67.3°C ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of 59.7g of cadmium from 25°C to 100°C ? The specific heat of cadmium is $0.231 \text{ J/g}^\circ\text{C}$. **ANS: 1034 J**

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