

# ACIDS, BASES AND SALTS

<b>Definition of an acid</b> An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	<b>What causes the acidic properties of acids?</b> The hydrogen ions ( $H^+$ ) cause the acidic properties and these are formed in the presence of water.
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<b>Another term to refer to an acid</b> An acid is called a <u>proton donor</u>	<b>Why an acid is also called a proton donor?</b> It's because an acid provides protons or hydrogen ions ( $H^+$ ) to other substances during the reaction.
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<b>Substance to which an acid provides protons</b> BASE	<b>A base is then called</b> PROTON ACCEPTOR	<b>Why the base is called a proton acceptor?</b> Because the base accepts hydrogen ions from acids	<b>Equation for the reaction between acid and base</b> $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid	HCl	MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for $H_2SO_4$ and nitrates for $HNO_3$
Sulphuric acid	$H_2SO_4$		
Nitric acid	$HNO_3$		

Other mineral acids known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	$H_2SO_3$	Derived from sulphites	Ethanoic acid ( $CH_3COOH$ )	CITRIC ACID from lemons
Carbonic acid	$H_2CO_3$	Derived from carbonates		TARTARIC ACID from grapes
Phosphoric acid	$H_3PO_4$	Derived from phosphates	Methanoic acid ( $HCOOH$ )	ACETIC ACID from vinegar
Nitrous acid	$HNO_2$	Derived from nitrites		LACTIC ACID from sour milk
				Hydrochloric acid from digestive juices

<b>Whenever an acid is dissolved in water, it produces</b> HYDROGEN IONS	<b>Term given to the number of hydrogen ions produced by one molecule of an acid</b> BASICITY OF AN ACID	<b>Definition of basicity of an acid</b> BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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# Guide Acids Bases And Salts

**Terry C. Jones**



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