Thermochemistry Worksheet Answers

Useful Values:

Heat of Fusion for Water: $\Delta H_{top} = 6 \text{ kJ/mol}$ Common error: not converting to moles Heat of Vaporization for Water: $\Delta H_{top} = 45 \text{ kJ/mol}$ Common error: same 1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K	
Water (ice)	2.05	0.490	Common error here:
Water (liquid)	4.18	1.00	using moles instead of grams.
Water (gas)	2.08	0.498	

How much energy does it take to warm the following substances the listed amount?

294 calories / 1230 Joules
600 calories / 2508 Joules
299 calories / 1249 Joules

To melt ice, one must bring ice to the freezing point (0° C), and then thaw it into water. How much heat would
it take to melt...

a. 30g of ice at 0° C?	10 kJ	(10,000 J)
b. 30g of ice at -40° C?	12.46 kJ	(12.460 J)

 To boil water, one must raise it to the boiling point (100° C), and then boil it into steam. How much heat would it take to boil...

a. 40g of water at 100° C?	100 kJ	(100,000 Joules)
b. 40g of water at 90° C?	101.7 kJ	(101,672 Joules)
c. 40g of water at 0° C?	116.7 kJ	(116,720 Joules)
d. 40g of ice at 0° C?	130.1 kJ	(130,053 Joules)
e. 40g of ice at -80° C?	136.6 kJ	(136,613 Joules)

4. If 700 calories of heat raise 400g of Phonium by 20° C, what is the specific heat of Phonium? 0.0875 cal/g/K

How much heat energy is stored in 2g of ice at -30° C? (remember: use Kelvin)
 238 cal / 996 J

How much heat energy is stored in 2g of Phonium at 70° C? (Assume Phonium is a solid at 70° C) 1672 cal.

7. If 100g of water at 10° C comes in contact with ______, what will the final temperature be? (Remember, you have the specific heat of Phonium from question 4.)

- a. 100g of water at 90° C
 50° C
- b. 100g of Phonium at 90° C (Hint: the water and Phonium end at the same temperature) 16.44° C Common error: 6.44 degrees is the temperature change, and the water gets warmer.
- c. 20g of Phonium at 90° C d. 1g of Phonium at 90° C 11.38° C 10.07° C

If 450g of Phakelite at 100° C is added to 1000g of water at 20° C, the temperature of the water increases by 5°.
 What is the specific heat of Phakalite? 0.148 cal/g/K

 If 450g of Nonsensium at 500° C is added to 1000g of water at 20° C, the temperature of the water increases by 25°. What is the specific heat of Nonsensium?
 0.122 cal/g/K

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