

# Lab Report: Gravimetric Analysis of an Unknown Sulfate

## Experimental Data

Unknown Sulfate ID Code: 20

1. Mass of empty 250-mL beaker	111.244 g
2. Mass of 250-mL beaker and unknown sulfate	112.227 g
3. Mass of unknown sulfate	0.343 g
4. Mass of empty crucible (without lid)	11.203 g
5. Mass of crucible (without lid) and barium sulfate	11.629 g
6. Mass of barium sulfate	0.426 g

## Calculations and Conclusions

- Calculate the mass of sulfate in the unknown sample. Clearly show each step of your calculation.

$$\begin{array}{c}
 \text{1 g BaSO}_4 \rightarrow \text{1 mol BaSO}_4 \rightarrow \text{1 mol SO}_4^{2-} \rightarrow \text{1 g SO}_4^{2-} \\
 \text{150.93 g/mol} \quad \quad \quad 233.37 \text{ g/mol} \quad \quad \quad 96.06 \text{ g/mol}
 \end{array}$$

$$0.426 \text{ g BaSO}_4 \times \frac{1 \text{ mol BaSO}_4}{233.37 \text{ g BaSO}_4} \times \frac{1 \text{ mol SO}_4^{2-}}{1 \text{ mol BaSO}_4} \times \frac{96.06 \text{ g SO}_4^{2-}}{1 \text{ mol SO}_4^{2-}} = 0.17524 \text{ g SO}_4^{2-}$$

$$\hookrightarrow \boxed{0.175 \text{ g SO}_4^{2-}}$$

- Calculate the percentage by mass of sulfate in the unknown sample. Show your work.  
mass of unknown sulfate: 0.343 g ; mass of sulfate: 0.175 g SO<sub>4</sub><sup>2-</sup>

$$\% \text{ SO}_4^{2-} = \frac{0.175 \text{ g}}{0.343 \text{ g}} \times 100 = 51.0204\%$$

$$\hookrightarrow \boxed{51.0\%}$$

- Calculate the percentage by mass of the metal in the unknown sample. Show your work.

$$100.0\% - 51.0\% = 49.0\%$$

# Gravimetric Analysis Lab Report Ucsd

**Jicheng Xie**



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