

$$M_{\text{AgNO}_3} = \frac{m_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{25.513 \text{ g} + 149.87 \frac{\text{g}}{\text{mL}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.569 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarities of the unknown AgNO_3 solution and the unknown NaCl solution. Round your data and show your calculations.

Unknown AgNO_3

Upon the addition of solid NaCl , solid AgCl was formed.

AgCl formed (g) = 2.32019 g

$$\text{mass}(\text{AgNO}_3) = \frac{M_r(\text{AgNO}_3)}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{149.87}{143.32} \frac{\text{g}}{\text{mol}} \times 2.32019 \text{ g} = \boxed{2.3500 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{m_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{2.3500 \text{ g} + 149.87 \frac{\text{g}}{\text{mL}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.16189 \frac{\text{mol}}{\text{L}}}$$

Unknown NaCl

Upon the addition of solid AgNO_3 , solid AgCl was formed.

AgCl formed (g) = 10.7907 g

$$\text{mass}(\text{NaCl}) = \frac{M_r(\text{NaCl})}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{58.443}{143.32} \frac{\text{g}}{\text{mol}} \times 10.7907 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{m_{\text{NaCl}}}{V_{\text{NaCl}}} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{\text{mL}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.75294 \frac{\text{mol}}{\text{L}}}$$

Conclusion questions

1. If you have a company that is using this method to create AgCl , knowing that AgNO_3 is an expensive chemical, what do you think would be a better limiting reagent to use, AgNO_3 or NaCl ? Why?

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Hugo de Lasa



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Table of Contents Limited Reagent Lab Key

1. Understanding the eBook Limited Reagent Lab Key
 - The Rise of Digital Reading Limited Reagent Lab Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limited Reagent Lab Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limited Reagent Lab Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limited Reagent Lab Key
 - Personalized Recommendations
 - Limited Reagent Lab Key User Reviews and Ratings
 - Limited Reagent Lab Key and Bestseller Lists
5. Accessing Limited Reagent Lab Key Free and Paid eBooks

- Limited Reagent Lab Key Public Domain eBooks
 - Limited Reagent Lab Key eBook Subscription Services
 - Limited Reagent Lab Key Budget-Friendly Options
6. Navigating Limited Reagent Lab Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limited Reagent Lab Key Compatibility with Devices
 - Limited Reagent Lab Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limited Reagent Lab Key
 - Highlighting and Note-Taking Limited Reagent Lab Key
 - Interactive Elements Limited Reagent Lab Key
 8. Staying Engaged with Limited Reagent Lab Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limited Reagent Lab Key
 9. Balancing eBooks and Physical Books Limited Reagent Lab Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limited Reagent Lab Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limited Reagent Lab Key
 - Setting Reading Goals Limited Reagent Lab Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limited Reagent Lab Key
 - Fact-Checking eBook Content of Limited Reagent Lab Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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