

### 1.1<sup>a</sup>. Concentration of a gas mixture.

A mixture of noble gases [helium (1), argon (2), krypton (3), and xenon (4)] is at a total pressure of 200 kPa and a temperature of 400 K. If the mixture has equal mole fractions of each of the gases, determine:

a) The composition of the mixture in terms of mass fractions.

Solution:

Basis: 100 kmole of the mixture

$$\begin{aligned}
 \text{ORIGIN} &:= 1 & \text{kmole} &:= 1000 \cdot \text{mole} \\
 \text{moles} &:= \begin{pmatrix} 25 \\ 25 \\ 25 \\ 25 \end{pmatrix} \cdot \text{kmole} & \text{MW} &:= \begin{pmatrix} 4 \\ 40 \\ 83.8 \\ 131.3 \end{pmatrix} \frac{\text{kg}}{\text{kmole}} & \text{mass} &:= \overrightarrow{(\text{moles} \cdot \text{MW})} \\
 \text{mass} &:= \begin{pmatrix} 100 \\ 1 \times 10^3 \\ 2.095 \times 10^3 \\ 3.283 \times 10^3 \end{pmatrix} \text{kg} & \text{total} &:= \sum_{i=1}^4 \text{mass}_i & \text{total} &:= 6.478 \times 10^3 \text{ kg} \\
 \text{mass\_fraction} &:= \frac{\text{mass}}{\text{total}} & \text{mass\_fraction} &:= \begin{pmatrix} 0.015 \\ 0.154 \\ 0.323 \\ 0.507 \end{pmatrix}
 \end{aligned}$$

b) The average molecular weight of the mixture.

Solution

$$\begin{aligned}
 \text{kmole} &:= 1000 \cdot \text{mole} \\
 M_{av} &:= \frac{\text{total}}{100 \cdot \text{kmole}} & M_{av} &= 0.065 \frac{\text{kg}}{\text{mol}}
 \end{aligned}$$

c) The total molar concentration.

Solution

$$\begin{aligned}
 P &:= 200 \cdot \text{kPa} & T &:= 400 \cdot \text{K} & \text{kPa} &:= 1000 \cdot \text{Pa} \\
 c &:= \frac{P}{R \cdot T} & R &:= 8.314 \cdot \text{kPa} \cdot \text{m}^3 \cdot \text{kmole}^{-1} \cdot \text{K}^{-1} \\
 c &:= 60.14 \frac{\text{mol}}{\text{m}^3}
 \end{aligned}$$

d) The mass density.

Solution

$$\begin{aligned}
 \rho &:= c \cdot M_{av} & \rho &= 3.896 \frac{\text{kg}}{\text{m}^3}
 \end{aligned}$$

# Mass Transfer Treybal Solution Manual

**R Bogdan**



## **Mass Transfer Treybal Solution Manual:**

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the costs of abatement justify the degree of abatement achieved This book is one of the volumes of the Handbook of Environmental Engineering series The principal intention of this series is to help readers formulate answers to the last two questions above The traditional approach of applying tried and true solutions to specific pollution problems has been a major contributing factor to the success of environmental engineering and has accounted in large measure for the establishment of a methodology of pollution control However the realization of the ever increasing complexity and interrelated nature of current environmental problems renders it imperative that intelligent planning of pollution abatement systems be undertaken

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